

Practice Problems Set #1

DIRECTIONS:

- YOU MUST ANSWER EVERY QUESTION IN ORDER TO GET ANY CREDIT!!!
- HIGHLIGHT EACH QUESTION NUMBER ON YOUR NOTEBOOK PAPER SO I CAN QUICKLY SEE THAT YOU HAVE DONE ALL THE PROBLEMS. IF I CAN'T FIND AN ANSWER, YOU WON'T GET CREDIT FOR ANY OF THE PROBLEMS!!!!
- HIGHLIGHT ANY QUESTION NUMBERS ON THIS PAGE THAT YOU WANT HELP WITH, HAVE QUESTIONS WITH, ETC!!!

Q #	QUESTION
1	Put 14850 in scientific notation.
2	Put 0.000387 in scientific notation.
3	Put 5.28×10^4 in standard notation
4	Put 8.75×10^{-3} in standard notation.
5	What is the metric prefix that represents 1000?
6	What is the metric prefix that represents 0.01?
7	How many millimeters are in 29.4 decameters?
8	How many kilometers are in 3405.2 centimeters?
9	Convert 4 mi/hr into m/s
10	Convert 19.2 mi/min into m/hr
11	Convert 52 m/s into mi/hr
12	What is the definition of a physical change?
13	What is the definition of a chemical change?
14	List 5 examples of physical changes.
15	List 5 examples of chemical changes.
16	What is the definition of atomic mass?
17	What does the atomic number tell you?
18	How many protons, neutrons, and electrons does an atom of silver have?
19	How many protons, neutrons, and electrons does each atom have? Cl Ba C Ne
20	What element has 12 protons?
21	How many neutrons does ^{14}C have? How many protons does it have? Electrons? Repeat with Carbon-12.
22	What is an electron orbital? How many electrons can fit in an orbital?
23	Sketch pictures of an "s" orbital and a "p" orbital.
24	How many e- can fit in a set of s orbitals? In a set of p orbitals? d orbitals? f orbitals?
25	Sketch an orbital diagram for sulfur and one for chlorine.
26	What element is represented by the e- configuration of: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^2$?
27	What element is represented by the electron configuration of: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
28	Write the electron configurations for H, He, K, Ca, Zn, I, Kr