

## Le Chatelier's Principle Worksheet



- 1) For the reaction below, which change would cause the equilibrium to shift to the right in this endothermic reaction?



- (a) Decrease the concentration of dihydrogen sulfide.
  - (b) Increase the pressure on the system.
  - (c) Increase the temperature of the system.
  - (d) Increase the concentration of carbon disulfide.
  - (e) Decrease the concentration of methane.
- 2) What would happen to the position of the equilibrium when the following changes are made to the equilibrium system below?



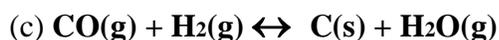
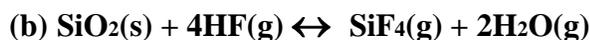
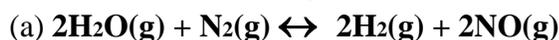
- (a) Sulfur dioxide is added to the system.
  - (b) Sulfur trioxide is removed from the system.
  - (c) Oxygen is added to the system.
- 3) What would happen to the position of the equilibrium when the following changes are made to the reaction below?



- (a) HgO is added to the system.
  - (b) The pressure on the system increases.
- 4) When the volume of the following mixture of gases is increased, what will be the effect on the equilibrium position?



- 5) Predict the effect of decreasing the volume of the container for each equilibrium.



- 6) Predict the effect of decreasing the temperature on the position of the following equilibria.

