

Periodic Trends WS

- 1) Where are the most active metals located? _____
- 2) Where are the most active non-metals located? _____
- 3) As you go from left to right across a period, does the atomic size decrease or increase. Why?

- 4) As you travel down a group, the atomic size, does decrease or increase. Why?

- 5) Is a negative ion larger or smaller than its parent atom? _____
- 6) Is a positive ion larger or smaller than its parent atom? _____
- 7) As you go from left to right across a period, does the first ionization energy generally decrease or increase? Why?

- 8) As you go down a group, does the first ionization energy generally decrease or increase? Why?

- 9) Where is the highest electronegativity found? _____
- 10) Where is the lowest electronegativity found? _____
- 11) Elements of Group 1A are called _____
- 12) Elements of Group 2A are called _____
- 13) Elements in the middle of the periodic table are called _____
- 14) Group 7A elements are called _____
- 15) Group 8A elements are called _____
- 16) As you go from left to right across the periodic table, do the elements go from (metals to nonmetals) or (nonmetals to metals)? _____
- 17) The most active element in Group 7A is _____
- 18) What sublevels (orbitals) are filling across the Transition Elements? _____

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19) Elements within a group have the same number of _____

20) The majority of elements in the periodic table are (metals / nonmetals). Circle one.

21) Elements in the periodic table are arranged according to their _____

22) ATOMIC RADIUS

For each of the following sets of atoms, rank the atoms from smallest to largest atomic radius.

- a. Li, C, F
- b. Li, Na, K
- c. Ge, P, O
- d. C, N, Al
- e. Al, Cl, Cu

23) IONIZATION ENERGY

For each of the following sets of atoms, rank them from lowest to highest ionization energy.

- a. Mg, Si, S
- b. Mg, Ca, Ba
- c. F, Cl, Br
- d. Ba, Cu, Ne
- e. Si, P, He

24) ELECTRONEGATIVITY

For each of the following sets of atoms, rank them from lowest to highest electronegativity.

- a. Li, C, N
- b. Ne, C, O
- c. Si, P, O
- d. Mg, K, P
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22) Elements within a group have the same number of _____

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