STUDY, STUDY, STUDY!
We use this ALL YEAR…

MEMORIZE!!!!!	
Name	**Formula**
Ammonium	NH_4^+
Silver	Ag^+
Cadmium	Cd^{2+}
Zinc	Zn^{2+}
Hydride	H^-
Hydroxide	OH^-
Chlorate	ClO_3^-
Chlorite	ClO_2^-
Nitrate	NO_3^-
Nitrite	NO_2^-
Carbonate	CO_3^{2-}
Peroxide	O_2^{2-}
Sulfate	SO_4^{2-}
Sulfite	SO_3^{2-}
Phosphate	PO_4^{3-}
Phosphite	PO_3^{3-}
From Periodic Table	**Transition metals**
Use periodic table Group 1A makes +1, Group 2A makes +2, etc…	All except Silver, Cadmium and Zinc need roman numerals. *Example: Fe^{+2} is Iron(II) and Fe^{+3} is Iron(III)*
Monoatomic ions	**Polyatomic ions**
Made of a single <u>type</u> of atom O_2^{2-}	Made of several <u>types</u> of atoms PO_4^{3-}
Cations	**Anions**
Lose electrons Make pos. charges	Gain electrons Make neg. charges

Will use, don't need to memorize	
Name	**Formula**
Hydronium	H_3O^+
Mercury (I)	Hg_2^{2+}
Mercury (II)	Hg^{2+}
Acetate	$C_2H_3O_2^-$
Bromate	BrO_3^-
Cyanide	CN^-
Thiocyanate	SCN^-
Hydrogen Carbonate (Bicarbonate)	HCO_3^-
Hydrogen Sulfate (Bisulfate)	HSO_4^-
Hydrogen Sulfite (Bisulfite)	HSO_3^-
Hypochlorite	ClO^-
Perchlorate	ClO_4^-
Iodate	IO_3^-
Permanganate	MnO_4^-
Chromate	CrO_4^{2-}
Dichromate	$Cr_2O_7^{2-}$
Hydrogen Phosphate (Biphosphate)	HPO_4^{2-}
Thiosulfate	$S_2O_3^{2-}$
Borate	BO_3^{3-}