# DAY TWO

# Polar

# Non-polar

# Jumpstart 2-A

1) Look at the molecules you drew for Jumsptart 1-A – redraw them and label them polar or non-polar
2) Remember how the lewis structure of water is bent? Is it polar or non-polar? If you drew it straight instead of bent why would that lead you to the wrong answer of polar or non-polar?

# Properties due to Intermolecular Forces

## PAGE 140

| TASK | Description | Observations |
|---|---|---|
| Penny | | |
| Mixing | | |
| Burettes | | |
| Hot Plates | | |

| TASK | Conclusions about forces |
|---|---|
| Penny | |
| Mixing | |
| Burettes | |
| Hot Plates | |

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**Target: I can make conclusions about the molecular forces present based on observable properties**

| K | C | Q |
|---|---|---|
| | | |

# Vocabulary

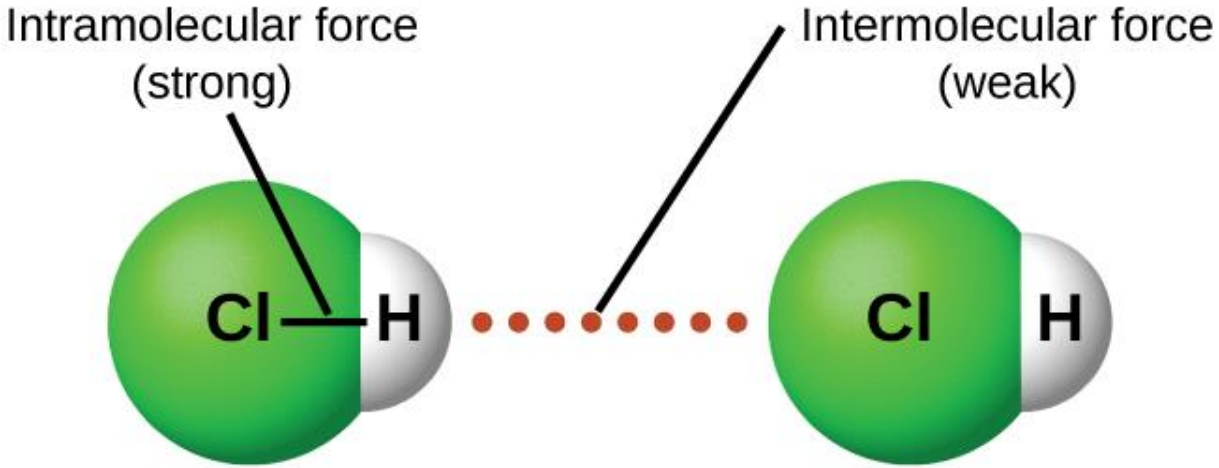
**INTRAmolecular Forces**

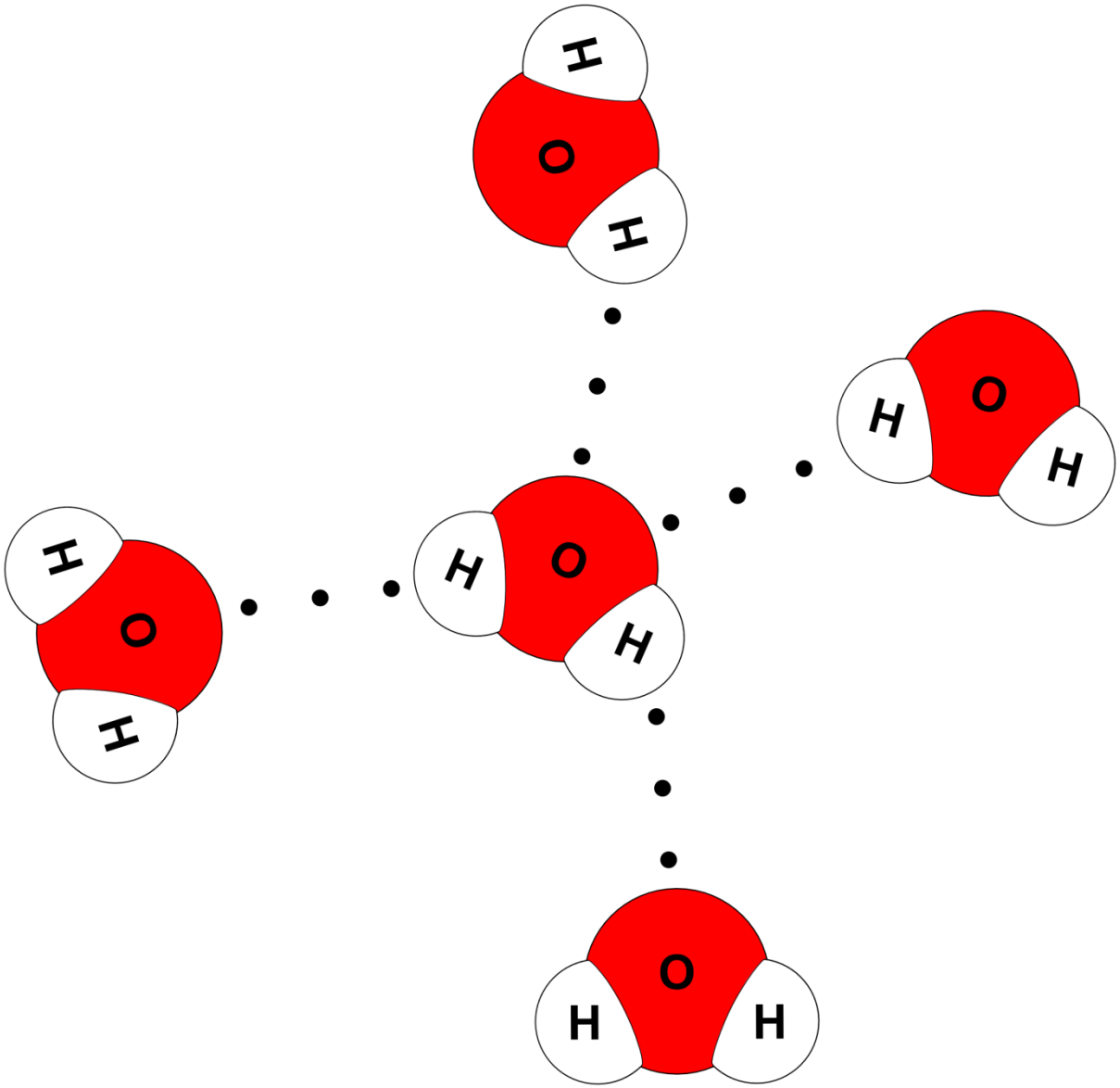
Forces holding together the atoms **INSIDE** a molecule or compound.

**Types:** Ionic forces, covalent forces

**INTERmolecular Forces**

Attractions or repulsions which act **between neighboring molecules**

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# Some properties that relate to intermolecular forces

<table>
<tr><td>Boiling point<br>Melting point<br>Viscosity<br>Surface tension</td><td colspan="2">↑forces = ↑</td></tr>
<tr><td rowspan="2">Mixing</td><td colspan="2">“Like dissolves like”</td></tr>
<tr><td>Polar<br>with polar</td><td>Non-polar<br>with non-polar</td></tr>
</table>

- https://www.youtube.com/watch?v=ohaI9yMZ8lc

| TASK | Description | Observations |
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| TASK | Conclusions about forces |
|---|---|
| Pennies | |
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**Target: I can make conclusions about the molecular forces present based on properties**

| K | C | Q |
|---|---|---|
| | | |