Balance the following chemical equations.

1. \( \_ \_ \text{CH}_4 + \_ \_ \text{O}_2 \rightarrow \_ \_ \text{CO}_2 + \_ \_ \text{H}_2\text{O} \)

2. \( \_ \_ \text{Na}^+ + \_ \_ \text{Cl}^- \rightarrow \_ \_ \text{NaCl} \)

3. \( \_ \_ \text{Al} + \_ \_ \text{O}_2 \rightarrow \_ \_ \text{Al}_2\text{O}_3 \)

4. \( \_ \_ \text{N}_2^+ + \_ \_ \text{H}_2 \rightarrow \_ \_ \text{NH}_3 \)

5. \( \_ \_ \text{CO} (\text{g}) + \_ \_ \text{H}_2(\text{g}) \rightarrow \_ \_ \text{C}_8\text{H}_{18}(\text{l}) + \_ \_ \text{H}_2\text{O} \)

6. \( \_ \_ \text{FeO}_3(\text{s}) + \_ \_ \text{CO} (\text{g}) \rightarrow \_ \_ \text{Fe}(\text{l}) + \_ \_ \text{CO}_2(\text{g}) \)

7. \( \_ \_ \text{H}_2\text{SO}_4 + \_ \_ \text{Pb} (\text{OH})_4 \rightarrow \_ \_ \text{Pb} (\text{SO}_4)_2 + \_ \_ \text{H}_2\text{O} \)

8. \( \_ \_ \text{Al} + \_ \_ \text{HCl} \rightarrow \_ \_ \text{AlCl}_3 + \_ \_ \text{H}_2 \)

9. \( \_ \_ \text{Ca}_3(\text{PO}_4)_2 + \_ \_ \text{H}_2\text{SO}_4 \rightarrow \_ \_ \text{CaSO}_4 + \_ \_ \text{Ca}(\text{H}_2\text{PO}_4)_2 \)

10. \( \_ \_ \text{H}_3\text{PO}_4 + \_ \_ \text{HCl} \rightarrow \_ \_ \text{PCl}_5 + \_ \_ \text{H}_2\text{O} \)