

**SPRING BENCHMARK #1 Review Problems – CHUNK #2**

<b>Q#</b>	<b><i>Balance the following equations using the smallest whole numbers possible.</i></b>	<b>Type</b>
1	$\text{___Mg(s)} + \text{___O}_2\text{(g)} \rightarrow \text{___MgO(s)}$	
2	$\text{___C}_8\text{H}_{18}\text{(g)} + \text{___O}_2\text{(g)} \rightarrow \text{___CO}_2\text{(g)} + \text{___H}_2\text{O(g)}$	
3	$\text{___Cu(s)} + \text{___H}_2\text{O(g)} \rightarrow \text{___H}_2\text{(g)} + \text{___Cu}_2\text{O}$	
4	$\text{___AgCl (aq)} + \text{___H}_2\text{S(aq)} \rightarrow \text{___Ag}_2\text{S(s)} + \text{___HCl(aq)}$	
5	$\text{___CaCO}_3\text{(s)} \rightarrow \text{___CaO(s)} + \text{___CO}_2\text{(g)}$	
6	$\text{___Cu(s)} + \text{___S}_8\text{(s)} \rightarrow \text{___CuS(s)}$	
7	$\text{___H}_2\text{S(aq)} + \text{___NaOH(aq)} \rightarrow \text{___H}_2\text{O(l)} + \text{___Na}_2\text{S(aq)}$	
8	$\text{___Al}_2\text{(SO}_4\text{)}_3 + \text{___Ca(OH)}_2 \rightarrow \text{___Al(OH)}_3 + \text{___CaSO}_4$	
9	$\text{___Al} + \text{___HCl} \rightarrow \text{___AlCl}_3 + \text{___H}_2$	
<b>Q#</b>	<b><i>Predict the products, balance the equation, then classify the type of reaction:</i></b>	<b>Type</b>
10	$\text{___Na} + \text{___FeBr}_3 \rightarrow$	
11	$\text{___NaOH} + \text{___H}_2\text{SO}_4 \rightarrow$	
12	$\text{___C}_2\text{H}_4\text{O}_2 + \text{___O}_2 \rightarrow$	
13	$\text{___NH}_3 + \text{___H}_2\text{O} \rightarrow$	
14	$\text{___PbSO}_4 + \text{___AgNO}_3 \rightarrow$	
15	$\text{___PBr}_3 \rightarrow$	
16	$\text{___HBr} + \text{___Fe} \rightarrow$	
17	$\text{___KMnO}_4 + \text{___ZnCl}_2 \rightarrow$	
18	$\text{___MnO}_2 + \text{___Sn(OH)}_4 \rightarrow$	
19	$\text{___O}_2 + \text{___C}_5\text{H}_{12}\text{O}_2 \rightarrow$	
20	$\text{___H}_2\text{O}_2 \rightarrow$	
21	$\text{___PtCl}_4 + \text{___Cl}_2 \rightarrow$	