

Precipitation Lab Follow-up Questions Answer the following questions in your notebook. Use complete sentences when possible. You may type your answers if you like. Include details, reference your lab data, and include diagrams/images/charts etc when applicable or useful.

- 1) Define precipitation and solubility
- 2) Each of the following equations shows what could happen when two solutions are mixed in a beaker.
Reaction 1: $\text{Pb}(\text{NO}_3)_2 + 2\text{NaCl} \rightarrow \text{PbCl}_2 + 2\text{NaNO}_3$
Reaction 2: $\text{FeSO}_4 + \text{Mg}(\text{NO}_3)_2 \rightarrow \text{MgSO}_4 + \text{Fe}(\text{NO}_3)_2$
 Identify which reaction would form a precipitate. Justify your answer by identifying the ions present in each solution before they are mixed, the precipitate that might form when they are mixed, explain why that precipitate forms.
- 3) The following pairs of solutions are mixed. Use the solubility rules to identify if a precipitate is formed

	Solutions that are mixed	Name of the PPT OR No Reaction
A	Silver nitrate + calcium chloride	
B	Potassium sulfate+ iron (II) nitrate	
C	Calcium nitrate+ sodium sulfate	

- 4) Write a balanced equation for the formation of ONE precipitate identified in question #2.
- 5) Complete the following word equations and include phases: Barium chloride + magnesium sulfate \rightarrow
- 6) Iron (II) nitrate solution is added to sodium hydroxide in a test tube.
 - A. Describe ONE observation that you would make as the reaction occurs.
 - B. Explain why your observation occurred.
 - C. Write a balanced equation for the rxn. Includes phases.
- 7) Be creative to summarize the concept of solubility and precipitation. Pictures, write a song, whatever. Be creative!

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