

**Predicting Products Practice**

You do NOT need to balancing the equations this time.

#	Type	Reaction
1		$\text{HgO} \rightarrow \text{Hg} + \text{O}_2$
2		$\text{NaCl} + \text{AgNO}_3 \rightarrow$
3		$\text{Mg} + \text{HCl} \rightarrow$
4		$\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow$
5		$\text{NaOH} + \text{HCl} \rightarrow$
6		$\text{CH}_4 + \text{O}_2 \rightarrow$
7		$\text{Al}_2(\text{SO}_4)_3 + \text{Ca}(\text{OH})_2 \rightarrow$
8		$\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
9		$\text{Cl}_2 + \text{NaBr} \rightarrow \text{NaCl} + \text{Br}_2$
10		$\text{Zn} + \text{CuSO}_4 \rightarrow$
11		$\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$
12		$\text{H}_2\text{O} + \text{Fe} \rightarrow$
13		$\text{Ca}(\text{OH})_2 + \text{HNO}_3 \rightarrow$
14	Synthesis	$\text{Na}_2\text{O} + \text{CO}_2 \rightarrow$
15		$\text{H}_2 + \text{N}_2 \rightarrow \text{NH}_3$

**Predicting Products Practice and Balancing Practice**

Predict the products and then balance.

#	Type of Rxn	Reaction and Products
1		$\text{HgO} + \text{Cl}_2 \rightarrow$
2	Synthesis	$\text{Na} + \text{Br}_2 \rightarrow$
3		$\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$
4		$\text{Ca}(\text{OH})_2 + \text{HNO}_3 \rightarrow$
5		$\text{Al}_2\text{O}_3 \rightarrow \text{Al} + \text{O}_2$
6		$\text{CuCl}_2 + \text{H}_2\text{S} \rightarrow$
7		$\text{NaOH} + \text{HBr} \rightarrow$
8		$\text{H}_2\text{O} + \text{Fe} \rightarrow$