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| Name: | Period: | Seat #: |
|--------------|----------------|----------------|

Directions: USE PENCIL!

- 1) Identify the type of reaction
- 2) Predict the products
- 3) Check to make sure the charges balance to neutral!
- 4) You do NOT need to balance the equation

Correcting:

- 1) The boxes on the right of the table will be highlighted to tell you what you need to fix for each problem.
- 2) Erase your old answer and write the new one in

| Q # | Reaction Type | Reactants | Products | Wrong Type | Wrong Products |
|-----|---------------|--|----------|------------|----------------|
| 1 | | $\text{Al} + \text{Pb}(\text{NO}_3)_2 \rightarrow$ | | | |
| 2 | | $\text{Al} + \text{CuCl}_2 \rightarrow$ | | | |
| 3 | | $\text{Fe} + \text{AgC}_2\text{H}_3\text{O}_2 \rightarrow$ | | | |
| 4 | | $\text{Fe}(\text{OH})_3 \rightarrow$ | | | |
| 5 | | $\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow$ | | | |
| 6 | | $\text{Ca}(\text{OH})_2 + \text{H}_3\text{PO}_4 \rightarrow$ | | | |
| 7 | | $\text{Cd}_3(\text{PO}_4)_2 + (\text{NH}_4)_2\text{S} \rightarrow$ | | | |
| 8 | | $\text{C}_{25}\text{H}_{52} + \text{O}_2 \rightarrow$ | | | |
| 9 | | $\text{Cl}_2 + \text{NaI} \rightarrow$ | | | |
| 10 | | $\text{C}_6\text{H}_6 + \text{O}_2 \rightarrow$ | | | |
| 11 | | $\text{MgCl}_2 + \text{O}_2 \rightarrow$ | | | |
| 12 | | $\text{Ni}(\text{ClO}_3)_2 \rightarrow$ | | | |
| 13 | | $\text{Na} + \text{O}_2 \rightarrow$ | | | |
| 14 | | $\text{BeO} + \text{CO}_2 \rightarrow$ | | | |