

Fall Final Exam Practice Problems - CHUNK #1 - Topics 1-7		
Topic #	Q #	Question
1	1	Who were the 8 scientists that were covered in class that contributed to the development of different atomic models? What did they each contribute?
	2	Sketch and name the seven atomic models that were covered in class.
	3	Describe Rutherford's Gold Foil experiment and what it proved.
	4	What does the Quantum model say about the nature of electrons in an atom?
2	5	Put 0.00345 in scientific notation
	6	Put 29800000 in scientific notation
	7	What is wrong with the following number that was supposed to be put in sci. notation? 24.6×10^3
	8	What is wrong with the following number that was supposed to be put in sci. notation? 0.54×10^2
3	9	Which prefix represents 1000?
	10	Which prefix represents 1/10?
	11	What is the "base unit" for length? For volume? For mass?
	12	What is the mnemonic we use for metric system?
4	13	How many centimeters are in 340.2 kilometers? (remember KHDBDCM)
	14	How many millimeters are in 29.4 meters?
	15	Convert 2.7 kg into grams.
	16	Convert 0.85 mg into decagrams.
5	17	What is the definition of atomic mass?
	18	What does the atomic number tell you?
	19	How many neutrons does an atom of silver have?
	20	How many protons, neutrons, and electrons does each atom have? Cl Ba C Ne
6	21	What is an isotope?
	22	What is the difference between Carbon-12, Carbon-13, and Carbon-14?
	23	How many protons, neutrons, electrons does Bromine-80 have compared to Bromine-83?
	24	Which version of Bromine above is the more common isotope? How do you know?
7	25	What is a mole?
	26	Why do we use "the mole" in chemistry class?
	27	Describe how to calculate molar mass for a molecule. Do you round your atomic masses or not? Why?
	28	What is the molar mass of Ag?
	29	What is the molar mass of $\text{Ca}(\text{OH})_2$
	30	What is the molar mass of K_2SO_4
	31	What is the molar mass of $(\text{NH}_4)_2\text{S}$,

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