

Dimensional Analysis Review

Perform the following conversions using the dimensional analysis technique. Fill in any portion that is missing

- 1) Convert 32 g to kg

$$\frac{32 \text{ g}}{1000 \text{ g}} \times \frac{1 \text{ kg}}{1000 \text{ g}} = \boxed{0.032 \text{ kg}}$$

- 2) Convert 12.5 mol to molecules

$$\frac{12.5 \text{ mol}}{1 \text{ mol}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} =$$

- 3) Convert 22.4 L to mL

$$\frac{22.4 \text{ L}}{1 \text{ L}} \times \frac{1000 \text{ mL}}{1 \text{ L}} =$$

- 4) Convert 5 m to cm

$$\frac{5 \text{ m}}{1 \text{ m}} \times \frac{100 \text{ cm}}{1 \text{ m}} =$$

- 5) Convert 17 in to ft

$$\frac{17 \text{ in}}{12 \text{ in}} \times \frac{1 \text{ ft}}{12 \text{ in}} =$$

- 6) Convert 1.3 g of H₂O to molecules

$$\frac{1.3 \text{ g}}{18 \text{ g}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} =$$

- 7) Convert 10.3 min to hr

$$\frac{10.3 \text{ min}}{60 \text{ min}} \times \frac{1 \text{ hr}}{60 \text{ min}} =$$

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$$\frac{10.3 \text{ min}}{60 \text{ min}} \times \frac{1 \text{ hr}}{60 \text{ min}} =$$

8) Convert 7.4×10^{24} molecules to mol

$$\frac{\text{molecules}}{\text{molecules}} \left| \frac{\text{mol}}{6.02 \times 10^{23}} \right| =$$

9) Convert 4315 mg of CO_2 to moles

$$\frac{\text{mg}}{\text{mg}} \left| \frac{\text{g}}{\text{g}} \right| =$$

10) Convert 0.82 mL to grams if density is 1.35 g/mL

11) Convert 0.5 m to cm

12) Convert 24 mi/hr to ft/min

$$\frac{24 \text{ mi}}{1 \text{ hr}} \left| \frac{\text{ft}}{\text{mi}} \right| \left| \frac{1 \text{ hr}}{60 \text{ min}} \right| =$$

13) Convert 16 mg/day to g/min

14) Convert 2210 mol of $\text{Fe}(\text{OH})_2$ to g

15) Convert 2.68×10^{15} molecules of H_2O to atoms

8) Convert 7.4×10^{24} molecules to mol

$$\frac{\text{molecules}}{\text{molecules}} \left| \frac{\text{mol}}{6.02 \times 10^{23}} \right| =$$

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