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|  | Atomic Number Practice |
| **Q #** | **Element Name**(don’t forget mass # at end!) | **Atomic Number** | **# of Protons** | **# of Neutrons** | **# of Electrons** | **Mass Number** |
| 1 | Carbon- |  |  |  |  | 12 |
| 2 |  | 8 |  | 8 |  |  |
| 3 | Hydrogen- |  |  |  |  | 1 |
| 4 |  |  | 6 |  |  | 14 |
| 5 | Hydrogen- |  |  | 2 |  |  |
| 6 | Nitrogen- |  |  |  |  | 14 |
| 7 |  |  |  | 1 |  | 2 |
| 8 |  | 92 |  | 146 |  |  |
| 9 | Cesium- |  |  | 82 |  |  |
| 10 |  | 11 |  | 12 |  |  |
| 11 |  |  | 47 |  |  | 108 |
| 12 | Tungsten- |  |  | 110 |  |  |
| 13 |  |  |  | 45 |  | 80 |
| 14 |  |  | 24 |  |  | 52 |
| 15 |  |  |  | 89 |  | 152 |
| 16 | Silver- |  |  |  |  | 107 |
| 17 |  | 76 |  | 114 |  |  |
| **Answer on Notebook Paper:** |
| 1. How are the atomic number and the number of protons related to each other?
 | 1. How do the number of protons, number of neutrons, and the mass number relate?
 | 1. What is the one thing that determines the identity (name) of an atom?
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| 2 |  | 8 |  | 8 |  |  |
| 3 | Hydrogen- |  |  |  |  | 1 |
| 4 |  |  | 6 |  |  | 14 |
| 5 | Hydrogen- |  |  | 2 |  |  |
| 6 | Nitrogen- |  |  |  |  | 14 |
| 7 |  |  |  | 1 |  | 2 |
| 8 |  | 92 |  | 146 |  |  |
| 9 | Cesium- |  |  | 82 |  |  |
| 10 |  | 11 |  | 12 |  |  |
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