0.02		per and Isoto	opes				
		in the box to con used more than	-	ving paragraph	about atomic	e mass.	/1
	number	standard	neutron(s)	proto	on(s) m	ass	
	The ele	ctron has very lit	tle mass compare	ed to the			_or
			. The mass of the	e atom depends of	on the nucleus	s and how	many
			and		it has.	The sum of	f the pro
ar	nd neutrons is	the mass		of an atom.	The number of	of neutrons	in an at
ca	an be found by	subtracting the a	atomic number fr	com the		nun	nber. Th
m	ass of the ator	n is so small that	there is a measu	rement called the	e atomic		
ur	nit.						
	nore than once many	in the box to con e. mixtures	protons	neutrons	between	num	
	ore than onc	e.			between		
	ore than once many element	e. mixtures	protons isotopes	neutrons six protons	between electrons	num	
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Directions: Complete the following chart and answer the questions below.

Remember:

Element Name – Mass # (Carbon-12) Atomic # = #Protons = #Electrons Atomic Mass = #Protons + #Neutrons (*rearrange it to find #protons, or #neutrons*)

Element Name	Atomic Number	# of Protons	# of Neutrons	# of Electrons	Mass Number
Carbon-					12
	8		8		
Hydrogen-					1
		6			14
Hydrogen-			2		
Nitrogen-					14
			1		2
	92		146		
Cesium-			82		
	11		12		
		47			108
Tungsten-			110		
			45		80
		24			52
			89		152
Silver-					107
	76		114		

- 1) How are the atomic number and the number of protons related to each other?
- 2) How do the number of protons, number of neutrons, and the mass number relate to each other?
- 3) What is the one thing that determines the identity of an atom (that is, whether it is an oxygen atom or a carbon atom, etc.)?