#1	$N_2O_{4\,(g)}$ <--------> $2NO_{2(g)}$ $\Delta H = +\ 92$ KJ			
The Stress	**Right or Left**	**$[N_2O_4]$**	**$[NO_2]$**	**Temperature**
$[N_2O_4]$ is increased	R	skip	INCR	DECR
$[NO_2]$ is increased	L	INCR	skip	INCR
Temp is increased	R	DECR	INCR	skip
$[N_2O_4]$ is decreased	L	skip	DECR	INCR
$[H_2]$ is decreased	NO AFFECT	NO AFFECT	NO AFFECT	NO AFFECT
$[NO_2]$ is decreased	R	DECR	skip	DECR
Temp is decreased	L	INCR	DECR	skip

#2	$4HCl_{(g)} + O_{2\,(g)}$ <--------> $2H_2O_{(g)} + 2Cl_{2\,(g)} + 98$ KJ				
The Stress	**Right or Left**	**[HCl]**	**$[O_2]$**	**$[H_2O]$**	**Temperature**
[HCl] is increased	R	skip	DECR	INCR	INCR
$[H_2O]$ is increased	L	INCR	INCR	skip	DECR
$[O_2]$ is increased	R	DECR	skip	INCR	INCR
Temp is increased	L	INCR	DECR	INCR	skip

#3	$CaCO_{3\,(s)} + 170$ KJ <----------> $CaO_{(s)} + CO_{2\,(g)}$ Reminder: Adding solids or liquids and removing solids or liquids does not shift the equilibrium. This is because you cannot change the concentration of a pure liquid or solid as they are 100% pure. It is only a concentration change that will change the # of collisions and hence shift the equilibrium.		
The Stress	**Right or Left**	**$[CO_2]$**	**Temperature**
$CaCO_3$ is added	NO CHANGE	NO CHANGE	NO CHANGE
CaO is added	NO CHANGE	NO CHANGE	NO CHANGE
CO_2 is added	L	skip	INCR
Temp is decreased	L	DECR	skip
A catalyst is added	NO CHANGE	NO CHANGE	NO CHANGE
$[CO_2]$ is decreased	R	skip	DECR
Temp is increased	R	INCR	skip
CaO is removed	NO CHANGE	NO CHANGE	NO CHANGE