

#1	$\text{N}_2\text{O}_4 (\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g}) \quad \Delta H = + 92 \text{ KJ}$				
The Stress	Right or Left	$[\text{N}_2\text{O}_4]$	$[\text{NO}_2]$	Temperature	
$[\text{N}_2\text{O}_4]$ is increased		skip			
$[\text{NO}_2]$ is increased			skip		
Temp is increased				skip	
$[\text{N}_2\text{O}_4]$ is decreased		skip			
$[\text{H}_2]$ is decreased					
$[\text{NO}_2]$ is decreased			skip		
Temp is decreased				skip	
#2	$4\text{HCl} (\text{g}) + \text{O}_2 (\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g}) + 2\text{Cl}_2 (\text{g}) + 98 \text{ KJ}$				
The Stress	Right or Left	$[\text{HCl}]$	$[\text{O}_2]$	$[\text{H}_2\text{O}]$	Temperature
$[\text{HCl}]$ is increased		skip			
$[\text{H}_2\text{O}]$ is increased				skip	
$[\text{O}_2]$ is increased			skip		
Temp is increased					skip
#3	$\text{CaCO}_3 (\text{s}) + 170 \text{ KJ} \rightleftharpoons \text{CaO} (\text{s}) + \text{CO}_2 (\text{g})$				
	Reminder: Adding solids or liquids and removing solids or liquids does not shift the equilibrium. This is because you cannot change the concentration of a pure liquid or solid as they are 100% pure. It is only a concentration change that will change the # of collisions and hence shift the equilibrium.				
The Stress	Right or Left	$[\text{CO}_2]$	Temperature		
CaCO_3 is added					
CaO is added					
CO_2 is added		skip			
Temp is decreased			skip		
A catalyst is added					
$[\text{CO}_2]$ is decreased		skip			
Temp is increased			skip		
CaO is removed					