#1	$N_2O_{4\,(g)}$ <--------> $2NO_{2(g)}$ $\Delta H = +\ 92$ KJ			
The Stress	**Right or Left**	**$[N_2O_4]$**	**$[NO_2]$**	**Temperature**
$[N_2O_4]$ is increased				
$[NO_2]$ is increased				
Temp is increased				
$[N_2O_4]$ is decreased				
$[H_2]$ is decreased				
$[NO_2]$ is decreased				
Temp is decreased				

#2	$4HCl_{(g)} + O_{2\,(g)}$ <--------> $2H_2O_{(g)} + 2Cl_{2\,(g)} + 98$ KJ				
The Stress	**Right or Left**	**[HCl]**	**$[O_2]$**	**$[H_2O]$**	**Temperature**
[HCl] is increased					
$[H_2O]$ is increased					
$[O_2]$ is increased					
Temp is increased					

#3	$CaCO_{3\,(s)} + 170$ KJ <----------> $CaO_{(s)} + CO_{2\,(g)}$ Reminder: Adding solids or liquids and removing solids or liquids does not shift the equilibrium. This is because you cannot change the concentration of a pure liquid or solid as they are 100% pure. It is only a concentration change that will change the # of collisions and hence shift the equilibrium.		
The Stress	**Right or Left**	**$[CO_2]$**	**Temperature**
$CaCO_3$ is added			
CaO is added			
CO_2 is added			
Temp is decreased			
A catalyst is added			
$[CO_2]$ is decreased			
Temp is increased			
CaO is removed			