

#1	$\text{N}_2\text{O}_4 (\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$				$\Delta H = + 92 \text{ KJ}$
The Stress	$[\text{N}_2\text{O}_4]$	$[\text{NO}_2]$	Right or Left	Reactants or Products	
$[\text{N}_2\text{O}_4]$ is increased					
$[\text{NO}_2]$ is increased					
Temp is increased					
$[\text{N}_2\text{O}_4]$ is decreased					
$[\text{H}_2]$ is decreased					
$[\text{NO}_2]$ is decreased					
Temp is decreased					
#2	$4\text{HCl} (\text{g}) + \text{O}_2 (\text{g}) \rightleftharpoons 2\text{H}_2\text{O} (\text{g}) + 2\text{Cl}_2 (\text{g}) + 98 \text{ KJ}$				
The Stress	$[\text{O}_2]$	$[\text{H}_2\text{O}]$	$[\text{HCl}]$	Right or Left	Reactants or Products
$[\text{HCl}]$ is increased					
$[\text{H}_2\text{O}]$ is increased					
$[\text{O}_2]$ is increased					
Temp is increased					
#3	$\text{CaCO}_3 (\text{s}) + 170 \text{ KJ} \rightleftharpoons \text{CaO} (\text{s}) + \text{CO}_2 (\text{g})$				
	Note: Adding solids or liquids and removing solids or liquids does not shift the equilibrium. This is because you cannot change the concentration of a pure liquid or solid as they are 100% pure. It is only a concentration change that will change the # of collisions and hence shift the equilibrium.				
The Stress	$[\text{CO}_2]$	Right or Left	Reactants or Products		
CaCO_3 is added					
CaO is added					
CO_2 is added					
Temp is decreased					
A catalyst is added					
$[\text{CO}_2]$ is decreased					
Temp is increased					
CaO is removed					