

1) Is this reaction exothermic or endothermic?



2) What is different about the equation below?



What do you think the difference means???

↔ Equilibrium! ↔

Some reactions can go forwards AND backwards



OR



↔ Equilibrium! ↔

- EQUILIBRIUM = the point at which the forward reaction is happening at the same **RATE** as the reverse reaction
- Are the CONCENTRATIONS of reactants and products the same?????
 - NO!!!!!! (well *maybe*, but it doesn't have to be!)

↔ Le Chatelier's principle ↔

If a stress is applied to a reaction at equilibrium
the reaction changes to relieve that stress

How do you “Stress” a reaction???

↔ Le Chatelier's principle ↔

Change:

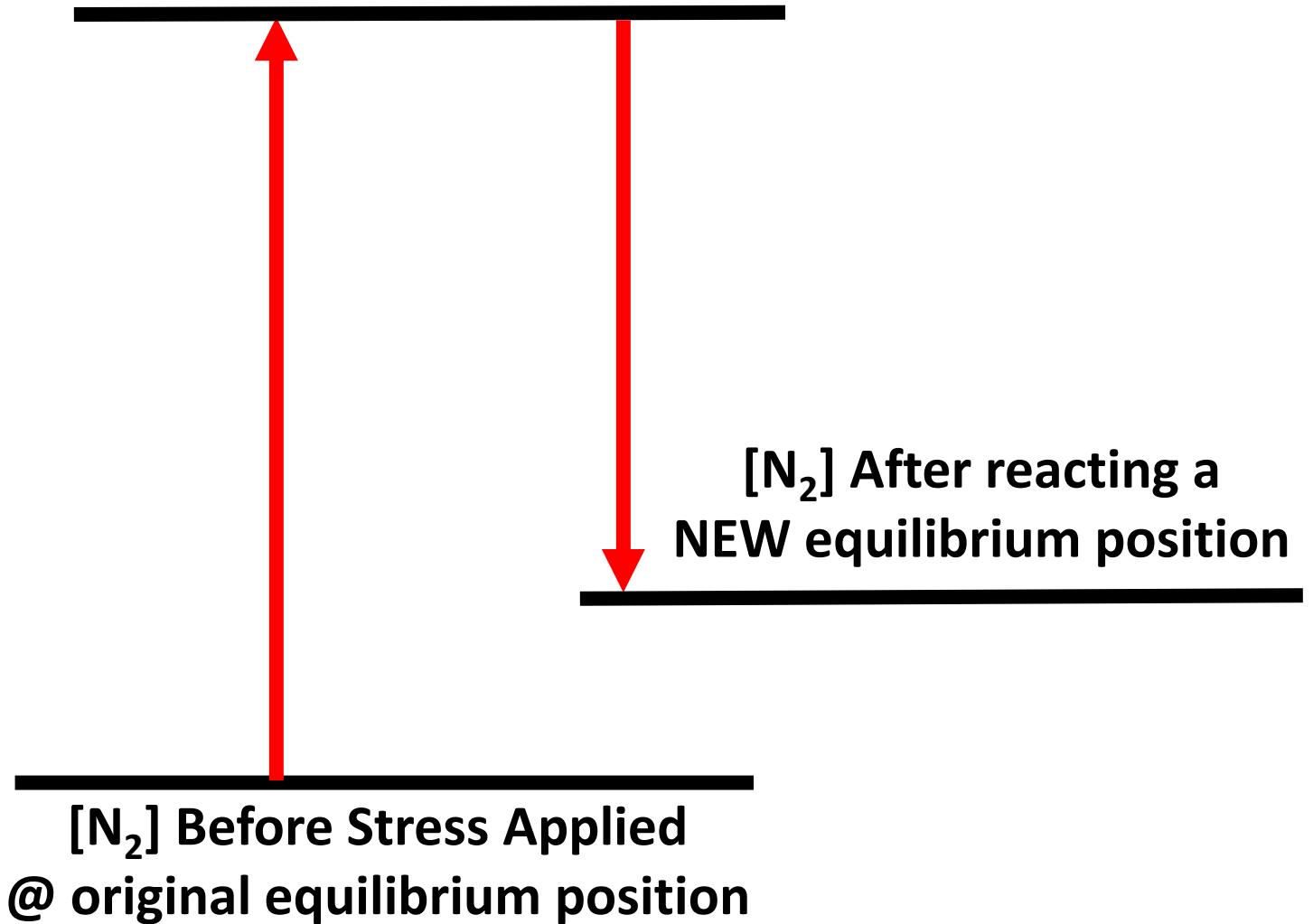
Temperature

Concentration

Pressure



**[N₂] During the Stress
no longer @ equilibrium**



↔ Let's Change the Concentration! ↔

