**Directions:** Complete the following chart by writing left, right, no change for equilibrium shift, and decrease, increase or no change for the concentration of reactant and product.

\_\_\_\_**N2 (g) + \_\_\_H2 (g) ⇔ \_\_\_NH3 (g) + 92.05 KJ**

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| **Stress** | **Equilibrium** | **[N2]** | **[H2]** | **[NH3]** |
| Add N2 | Right | Increase | Decreases | Increases |
| Add H2 |  |  |  |  |
| Add NH3 |  |  |  |  |
| Remove N2 |  |  |  |  |
| Remove H2 |  |  |  |  |
| Remove NH3 |  |  |  |  |
| Increases Temperature |  |  |  |  |
| Decreases Temperature |  |  |  |  |
| Increase Pressure |  |  |  |  |
| Decrease Pressure |  |  |  |  |