

Specific Heat of Brass Lab Questions

- 1) Why do we use the ending temperature of the water as the ending temperature of the brass?
- 2) Why can we assume the starting temperature of the brass is 100° ?
- 3) How do you convert from Q_{water} into Q_{brass} ?
- 4) Draw a diagram that explains your answer to question #3
- 5) What are three sources of error that probably happened in your lab? Do not just list anything that could have happened to someone doing the lab – they need to be things that probably happened in YOUR lab.
- 6) Why is Q of water a positive value? Does that make it endothermic or exothermic?
- 7) Why is Q of brass a negative value? Does that make it endothermic or exothermic?
- 8) Did the size and shape of the brass make a difference in the specific heat value? Why or why not?
- 9) What is the “key” to calorimetry? What is the conceptual part of the math that makes it a calorimetry lab?
- 10) How would your data change if we had put the brass in the freezer and put hot water into the calorimeter? How would the math change? How would your answer change?