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| Thermochemistry Webquest |
| **Part 1 - Vocabulary** | Click on the first letter of the word you are looking for the definition of: ttps://goo.gl/o93Uk7 |
| **1** | Kinetic Energy |  |
| **2** | Potential Energy |  |
| **3** | Endothermic |  |
| **4** | Exothermic |  |
| **5** | Activation Energy |  |
| **6** | Heat of Reaction |  |
| **7** | Enthalpy |  |
| **8** | Entropy |  |
| **9** | Specific Heat |  |
| **Part 2 - Watch a Video** | Watch the Video at the link below. Then answer the following questions. ttp://goo.gl/LHZTPO |
| **1** | What does it mean to have thermoenergy |  |
| **2** | What temperature would something need to be, in order to not have thermoenergy.  |  |
| **3** | What is the difference between potential and kinetic energy? |  |
| **4** | What is thermodynamics? |  |
| **5** | What is the equation for change of energy to a system? |  |
| **6** | What is the difference between an exothermic and endothermic reaction? |  |
| **7** | Do you think the ΔE for an exothermic reaction would be positive or negative? |  |
| **Part 3 - Specific Heat** | Use the following link to answer the questions below: https://goo.gl/IEbRz2 |
| **1** | Write the formula that can be used to calculate the heat energy being absorbed or released in a system |  |
| **2** | Describe each part of the equation: heat gained or lost = Mass x Change in Temperature x Specific Heat | Mass | Change in Temp | Specific Heat |
| **3** | Using the Table link at the bottom of the equation box, what solid substance has the highest specific heat? |  |