Periodic Trends Lab – Post Lab Questions

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| 1. Which metal reacted faster with water?
 | 1. Which metal reacted faster with acid?
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| 1. Make a statement about the trends in reactivity as you move down the column of alkaline earth metals.
 | 1. Predict the reactivity of strontium and barium, based on your activity in this lab.
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| 1. If sufficient radium could be gathered for a test, predict it reactivity with water and hydrochloric acid. Explain.
 | 1. Why would it be dangerous to handle even a small amount of radium? Your answer should be related to this lab and the concept of reactivity NOT radioactivity.
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| 1. Group VIIA (The halogens – nonmetallic elements) reactivity *decreases* as the atomic number *increases*. Why do you think this group of elements is opposite Group IA? Explain. (HINT – think about atomic structure, valence electrons, electronegativity)
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| 1. Was this a *quantitative* or *qualitative* lab? Why?
 | 1. Brainstorm a way that you could add a quantitative aspect to this lab.
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