## **Lewis Structures of Atoms, Ions and Ionic Compounds WS**

#	Question	Answer	
1	What is the octet rule? Explain its role in bonding between atoms		
2	Indicate how many electrons must be gained or lost by each of the following atoms to achieve a stable electron configuration (3 lost, 2 gained, etc)	a) Sr b) Sb c) Si	d) S e) Se f) Xe
3	Which of the following pairs of elements will <u>not</u> form ionic compounds? Explain why or why not for each pair.	Sulfur and Xenon Strontium and Sulfur	Sodium and Calcium  Selenium and Chlorine
4	Draw the Lewis Structure for the a) K	e following Ions or Comp	pounds
	b) Ba <sup>2+</sup>	g) BaF <sub>2</sub>	
	c) C	h) C <sup>4-</sup>	
	d) MgO	i) S <sup>2-</sup>	
	e) Br	j) Na₂S	

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	not for each pair.	Chlorine		
	Draw the Lewis Structure for the following Ions or Compounds			
4	a) K	f) K <sup>+</sup>		
	b) Ba <sup>2+</sup>	g) BaF <sub>2</sub>		
	c) C	h) C <sup>4-</sup>		
	d) MgO	i) S <sup>2-</sup>		
	e) Br	j) Na <sub>2</sub> S		