|  |  |  |
| --- | --- | --- |
| Ionic Bonding Worksheet | | |
| **#** | **Question** | **Answer** |
| **1** | What are the electrons in the outer shell of an atom that form the chemical bonds called? |  |
| **2** | What are the three main types of bonds? |  |
| **3** | Using the terms: metal and nonmetal, describe what types of atoms make up each of the three types of  bonds mentioned above. |  |
| **4** | What is the name of the type of chemical bond that is formed when two electrons are shared? |  |
| **5** | What is different between an ionic bond and the bond in the above question? |  |
| **6** | How many valance electrons are there in carbon? Where do you look on the periodic table to find out? |  |
| **7** | Label the following as ionic, or covalent, compounds/molecules | |
| 1. NaCl 2. CO2 3. K3P 4. HBr | 1. CCl4 2. AlBr3 3. H2O |
| **8** | How many Carbon atoms are in CO2 ?  How many oxygen atoms are in CO2 ? |  |
| **9** | How many Nitrogen atoms are in (NH4)2S ?  How many Hydrogen atoms? How many Sulfur atoms |  |
| **10** | Explain how to name binary ionic compounds |  |
| **11** | Explain how to name polyatomic ionic compounds |  |
| **12** | Name the following ionic compounds: | |
| 1. LiOH 2. Na2SO4 3. SbCl3 4. Al(OH)3 5. Sb(NO3)3 | 1. Al2(SO4)3 2. HgO 3. Fe2S3 4. Pb(NO3)2 5. K2SO3 |

|  |  |  |
| --- | --- | --- |
| Ionic Bonding Worksheet | | |
| **#** | **Question** | **Answer** |
| **1** | What are the electrons in the outer shell of an atom that form the chemical bonds called? |  |
| **2** | What are the three main types of bonds? |  |
| **3** | Using the terms: metal and nonmetal, describe what types of atoms make up each of the three types of  bonds mentioned above. |  |
| **4** | What is the name of the type of chemical bond that is formed when two electrons are shared? |  |
| **5** | What is different between an ionic bond and the bond in the above question? |  |
| **6** | How many valance electrons are there in carbon? Where do you look on the periodic table to find out? |  |
| **7** | Label the following as ionic, or covalent, compounds/molecules | |
| 1. NaCl 2. CO2 3. K3P 4. HBr | 1. CCl4 2. AlBr3 3. H2O |
| **8** | How many Carbon atoms are in CO2 ?  How many oxygen atoms are in CO2 ? |  |
| **9** | How many Nitrogen atoms are in (NH4)2S ?  How many Hydrogen atoms? How many Sulfur atoms |  |
| **10** | Explain how to name binary ionic compounds |  |
| **11** | Explain how to name polyatomic ionic compounds |  |
| **12** | Name the following ionic compounds: | |
| 1. LiOH 2. Na2SO4 3. SbCl3 4. Al(OH)3 5. Sb(NO3)3 | 1. Al2(SO4)3 2. HgO 3. Fe2S3 4. Pb(NO3)2 5. K2SO3 |