

Ionic Bonding Worksheet		
#	Question	Answer
1	What are the electrons in the outer shell of an atom that form the chemical bonds called?	
2	What are the three main types of bonds?	
3	Using the terms: metal and nonmetal, describe what types of atoms make up each of the three types of bonds mentioned above.	
4	What is the name of the type of chemical bond that is formed when two electrons are shared?	
5	What is different between an ionic bond and the bond in the above question?	
6	How many valence electrons are there in carbon? Where do you look on the periodic table to find out?	
Label the following as ionic, or covalent, compounds/molecules		
7	a) NaCl b) CO <sub>2</sub> c) K <sub>3</sub> P d) HBr	e) CCl <sub>4</sub> f) AlBr <sub>3</sub> g) H <sub>2</sub> O
8	How many Carbon atoms are in CO <sub>2</sub> ? How many oxygen atoms are in CO <sub>2</sub> ?	
9	How many Nitrogen atoms are in (NH <sub>4</sub> ) <sub>2</sub> S? How many Hydrogen atoms? How many Sulfur atoms	
10	Explain how to name binary ionic compounds	
11	Explain how to name polyatomic ionic compounds	
Name the following ionic compounds:		
12	a) LiOH b) Na <sub>2</sub> SO <sub>4</sub> c) SbCl <sub>3</sub> d) Al(OH) <sub>3</sub> e) Sb(NO <sub>3</sub> ) <sub>3</sub>	f) Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> g) HgO h) Fe <sub>2</sub> S <sub>3</sub> i) Pb(NO <sub>3</sub> ) <sub>2</sub> j) K <sub>2</sub> SO <sub>3</sub>

Ionic Bonding Worksheet		
#	Question	Answer
1	What are the electrons in the outer shell of an atom that form the chemical bonds called?	
2	What are the three main types of bonds?	
3	Using the terms: metal and nonmetal, describe what types of atoms make up each of the three types of bonds mentioned above.	
4	What is the name of the type of chemical bond that is formed when two electrons are shared?	
5	What is different between an ionic bond and the bond in the above question?	
6	How many valence electrons are there in carbon? Where do you look on the periodic table to find out?	
Label the following as ionic, or covalent, compounds/molecules		
7	h) NaCl i) CO <sub>2</sub> j) K <sub>3</sub> P k) HBr	l) CCl <sub>4</sub> m) AlBr <sub>3</sub> n) H <sub>2</sub> O
8	How many Carbon atoms are in CO <sub>2</sub> ? How many oxygen atoms are in CO <sub>2</sub> ?	
9	How many Nitrogen atoms are in (NH <sub>4</sub> ) <sub>2</sub> S? How many Hydrogen atoms? How many Sulfur atoms	
10	Explain how to name binary ionic compounds	
11	Explain how to name polyatomic ionic compounds	
Name the following ionic compounds:		
12	k) LiOH l) Na <sub>2</sub> SO <sub>4</sub> m) SbCl <sub>3</sub> n) Al(OH) <sub>3</sub> o) Sb(NO <sub>3</sub> ) <sub>3</sub>	p) Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> q) HgO r) Fe <sub>2</sub> S <sub>3</sub> s) Pb(NO <sub>3</sub> ) <sub>2</sub> t) K <sub>2</sub> SO <sub>3</sub>