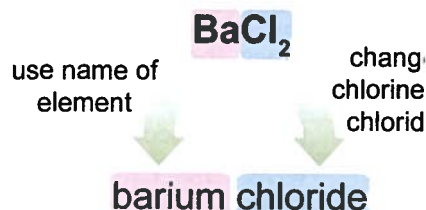
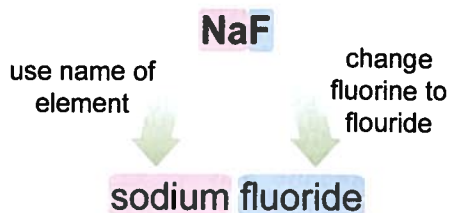


Writing names for ionic compounds

Naming binary ionic compounds

When naming an ionic compound, you just write the names of the ions as they are in the formula. For simple binary ionic compounds (ionic formulas from two simple ions), like NaCl, the first part of the name is just the name of the element, and the second part of the name is a modified version of the name of the element. Typically, you use a standard ending and add -ide.



Names for polyatomic ions are found on the common ion table

For polyatomic ions you will need to get the name from the table below. Others are done the same way as it is for binary ionic compounds. Write the name of the ion in the order that they appear in the formula.

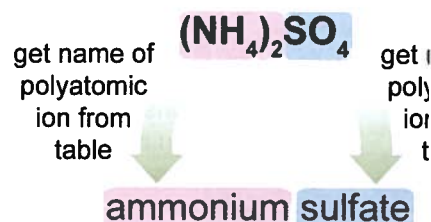
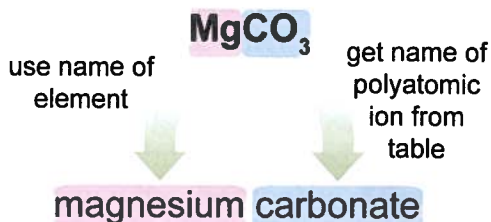


TABLE 8.1. Common Ions

positive ions		negative ions	
Aluminum Al ³⁺	Lead(II) Pb ²⁺	Acetate C ₂ H ₃ O ₂ ¹⁻	Hydrogen sulfate (bisulfate) HSO ₄ ¹⁻
Ammonium NH ₄ ¹⁺	Lead(IV) Pb ⁴⁺	Chloride Cl ¹⁻	Hydroxide OH ¹⁻
Barium Ba ²⁺	Magnesium Mg ²⁺	Carbonate CO ₃ ²⁻	Nitrate NO ₃ ¹⁻
Copper(I) Cu ¹⁺	Mercury(I) Hg ₂ ¹⁺	Chromate CrO ₄ ²⁻	Nitrite NO ₂ ¹⁻
Copper(II) Cu ²⁺	Mercury(II) Hg ²⁺	Chlorate ClO ₃ ¹⁻	Oxide O ²⁻
Calcium Ca ²⁺	Potassium K ¹⁺	Chlorite ClO ₂ ¹⁻	Peroxide O ₂ ²⁻
Chromium(II) Cr ²⁺	Silver Ag ¹⁺	Cyanide CN ¹⁻	Phosphate PO ₄ ³⁻
Chromium(III) Cr ³⁺	Sodium Na ¹⁺	Dichromate Cr ₂ O ₇ ²⁻	Sulfate SO ₄ ²⁻
Hydrogen* H ¹⁺	Tin(II) Sn ²⁺	Fluoride F ¹⁻	Sulfite SO ₃ ²⁻
Iron(II) Fe ²⁺	Tin(IV) Sn ⁴⁺	Hydrogen carbonate (bicarbonate) HCO ₃ ¹⁻	Sulfide S ₂ ²⁻
Iron(III) Fe ³⁺	Zinc Zn ²⁺		

*Hydrogen ions rarely exist by themselves. Often they combine with water to form hydronium ions: H₃O¹⁺.