Regular Chem Winter Break Reminders and Suggestions:

When we return from Winter Break we will be starting a new chapter called “Intermolecular Forces.” This chapter starts to look at how molecules interact with each other when next to each other. After that we will start our “Reactions” chapter which is where we learn how molecules react to form new molecules. These two chapters combine almost every topic we have learned during 1st semester.

There is no official homework over Winter Break, however, please make sure that you do not forget the following topics while on vacation! If you struggled with these topics during 1st semester please spend some time reviewing the topics. We want to make sure that everyone comes back from Winter Break ready to start 2nd semester off in a strong way!

Included in this packet is a chart of pages where you learned the topics, pages where you have practice problems we did during the year, and a small practice test of some examples of the types of things we need to make sure we don’t forget how to do. Please realize that this practice test is not required, and it does not show every single possible thing you need to remember from 1st semester, it is just a couple examples to remind you.

We will keep using the same Interactive Notebook 2nd semester so do not lose your notebook or get a new one. The gradebook starts over 2nd semester so everyone gets to start fresh and work towards completing all their work, doing well on quizzes and tests, etc.

If you have questions please email me – I will not be checking email daily, but I will check it occasionally over vacation. Thank you, and have a fabulous Winter Break!

Mrs. Farmer

Review and/or practice the following topics:

1. Study your ions!

* ***There will be an ion quiz the week we return!***
* The day is unannounced, but it will be during the first week.
* Remember to know the ones on your green ion sheet, but also any atoms from the periodic table s, p, d block that follow the pattern of the groups:   
  1A = +1, 2A = +2, 3A = +3, 4A = +/-4, 5A = -3, 6A = -2, 7A = -1, 8A = no charge

1. Periodic trend for electronegativity

* Identify which atom is more electronegative

1. Types of bonds

* Identify if a molecule is ionic or covalent

1. Writing formulas

* Crossing over to make neutral ionic compounds
* Using prefixes to write covalent molecules

1. Naming ionic and covalent molecules

* Remember, there are two different ways to name things – one for ionic, one for covalent

1. Lewis Structures

* The “Intermolecular Forces” chapter looks at how symmetrical & unsymmetrical molecules behave. Without a correct Lewis Structure we won’t know if it’s symmetrical or not!

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| Notebook Pages to Review - www.mychemistryclass.net | | |
| Topic | Notes Page(s) | Worksheet Page(s) |
| Ions | 43 (ions in general) ,101 | 100 |
| Electronegativity | 85 | 84, 86, 87 |
| Types of Bonds | 103 | 102, 104 |
| Writing/Naming Formulas | 107, 111 | 106, 108, 109, 110, 111, 112, 113 |
| Lewis Structures | 117, 119, 121 | 116, 122, 123, 124, 125 |

\*Remember – You have QR codes in your notebook of videos that review the topics, the class website has a “Resources” tab that has links to other websites and other practice, and you have the entire internet at your fingertips too! ☺

Practice Test for Jogging Your Memory: