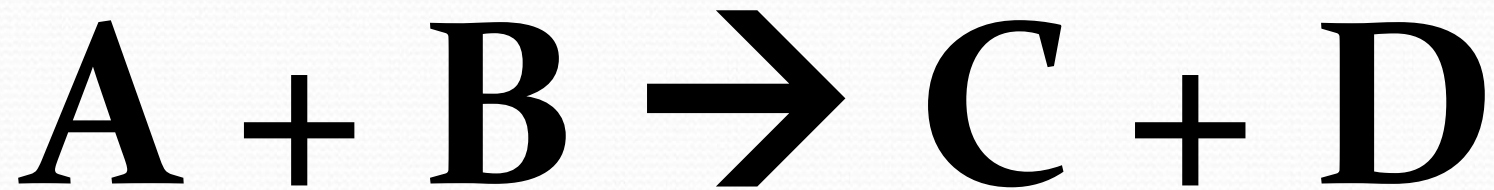


Balancing Chemical Equations



The Basics



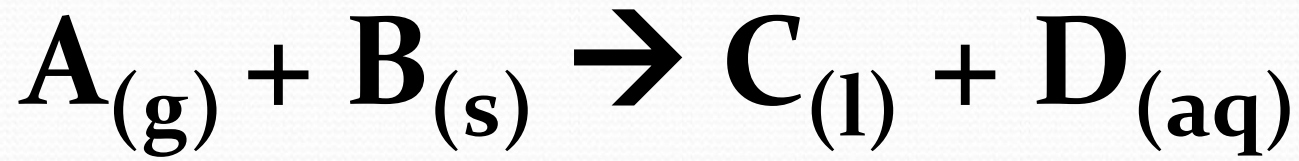
Reactants

(starting materials)

Products

(ending materials)

Phases



g = gas

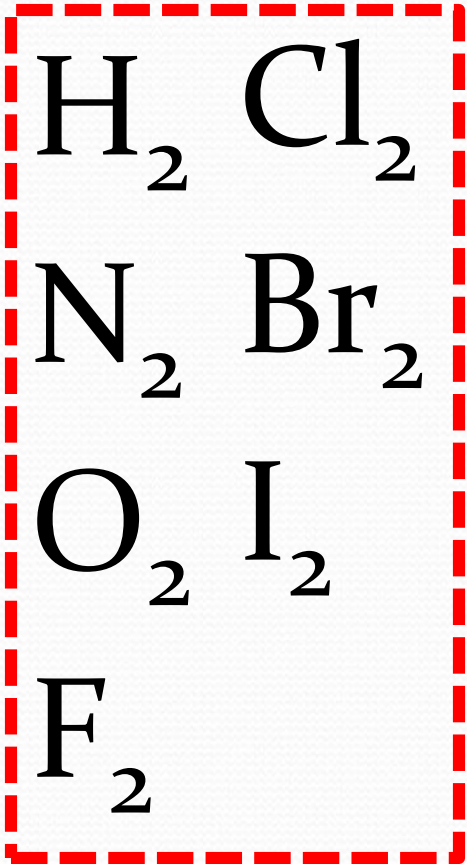
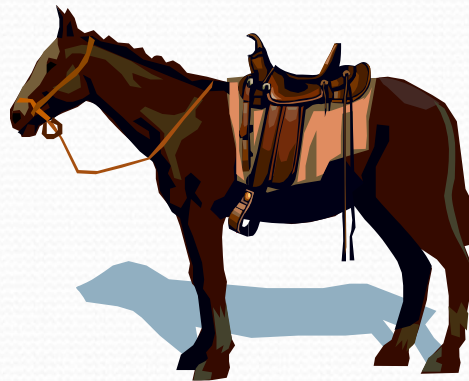
s = solid

l = liquid

aq = “aqueous” – ions in water

Diatomic Gases

Horses Need
Oats For Clear
Brown “Eyes”



Writing Equations

Word Equations

Written with the names of the compounds

*hydrogen gas and chlorine gas
combine to form hydrogen chloride gas*

Skeleton Equations

Written with formulas



Balancing Equations

Law of Conservation of Matter

Matter is not created or destroyed

of atoms for each element before and after the reaction must be equal.



- **Reactants:** 2 H & 2 Cl
- **Products** 2 H & 2 Cl
- The two sides are balanced!