

## SPRING BENCHMARK #1 Review Problems – CHUNK #2

Q#	Balance the following equations using the smallest whole numbers possible.	Type
1	$\text{___ Mg(s)} + \text{___ O}_2\text{(g)} \rightarrow \text{___ MgO(s)}$	
2	$\text{___ C}_8\text{H}_{18}\text{(g)} + \text{___ O}_2\text{(g)} \rightarrow \text{___ CO}_2\text{(g)} + \text{___ H}_2\text{O(g)}$	
3	$\text{___ Cu(s)} + \text{___ H}_2\text{O(g)} \rightarrow \text{___ H}_2\text{(g)} + \text{___ Cu}_2\text{O}$	
4	$\text{___ AgCl (aq)} + \text{___ H}_2\text{S(aq)} \rightarrow \text{___ Ag}_2\text{S(s)} + \text{___ HCl(aq)}$	
5	$\text{___ CaCO}_3\text{(s)} \rightarrow \text{___ CaO(s)} + \text{___ CO}_2\text{(g)}$	
6	$\text{___ Cu(s)} + \text{___ S}_8\text{(s)} \rightarrow \text{___ CuS(s)}$	
7	$\text{___ H}_2\text{S(aq)} + \text{___ NaOH(aq)} \rightarrow \text{___ H}_2\text{O(l)} + \text{___ Na}_2\text{S(aq)}$	
8	$\text{___ Al}_2\text{(SO}_4\text{)}_3 + \text{___ Ca(OH)}_2 \rightarrow \text{___ Al(OH)}_3 + \text{___ CaSO}_4$	
9	$\text{___ Al} + \text{___ HCl} \rightarrow \text{___ AlCl}_3 + \text{___ H}_2$	
Q#	Predict the products, balance the equation, then classify the type of reaction:	Type
10	$\text{___ Na} + \text{___ FeBr}_3 \rightarrow$	
11	$\text{___ NaOH} + \text{___ H}_2\text{SO}_4 \rightarrow$	
12	$\text{___ C}_2\text{H}_4\text{O}_2 + \text{___ O}_2 \rightarrow$	
13	$\text{___ NH}_3 + \text{___ H}_2\text{O} \rightarrow$	
14	$\text{___ PbSO}_4 + \text{___ AgNO}_3 \rightarrow$	
15	$\text{___ PBr}_3 \rightarrow$	
16	$\text{___ HBr} + \text{___ Fe} \rightarrow$	
17	$\text{___ KMnO}_4 + \text{___ ZnCl}_2 \rightarrow$	
18	$\text{___ MnO}_2 + \text{___ Sn(OH)}_4 \rightarrow$	
19	$\text{___ O}_2 + \text{___ C}_5\text{H}_{12}\text{O}_2 \rightarrow$	
20	$\text{___ H}_2\text{O}_2 \rightarrow$	
21	$\text{___ PtCl}_4 + \text{___ Cl}_2 \rightarrow$	