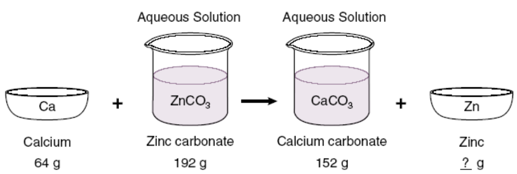
<u>Signs of a Reaction and Conservation of Matter Lab Follow-up Questions</u> *Answer the following questions in your notebook. Use complete sentences, use details, draw pictures if you need/want to. You may type your answers if you wish.*

1) List the things you observed during the lab that were signs a reaction took place.
2) Describe how this lab related to the Law of Conservation of Matter. Make sure to include details and observations from your lab.
3) Define exothermic and endothermic.
4) Describe how those terms relate to the lab. Make sure to use observational data from your lab.
5) Did the lab work perfectly? Were there deviations from perfect results? What could have caused these results?
6) In the following reaction, $2Na_3N$ decomposes to form 2Na plus $3N_2$. If 500 g of NaN_3 decomposes to form 323.20 g of N_2, How much Na is produced?
7) Which of the following best illustrates the Law of Conservation of Mass? Explain why.
   A. $H_2O_2 \rightarrow H_2O + O_2$
   B. $Na + CuS \rightarrow Na_2S + Cu$
   C. $K + AgCl \rightarrow KCl + Ag$
   D. $NaOH + 2\,HCl \rightarrow NaCl + H_2O$
8) Use the diagram below. According to the law of conservation of mass, how much zinc was present in the zinc carbonate? Explain or diagram how you figured this out.



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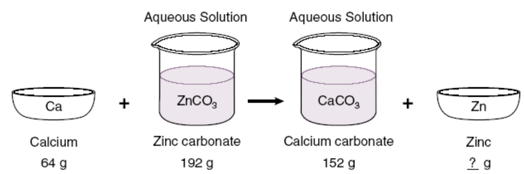

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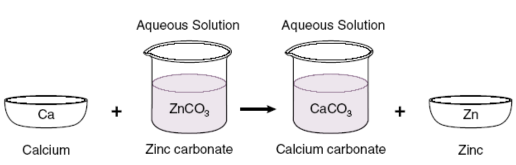

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