

# Types of Reactions

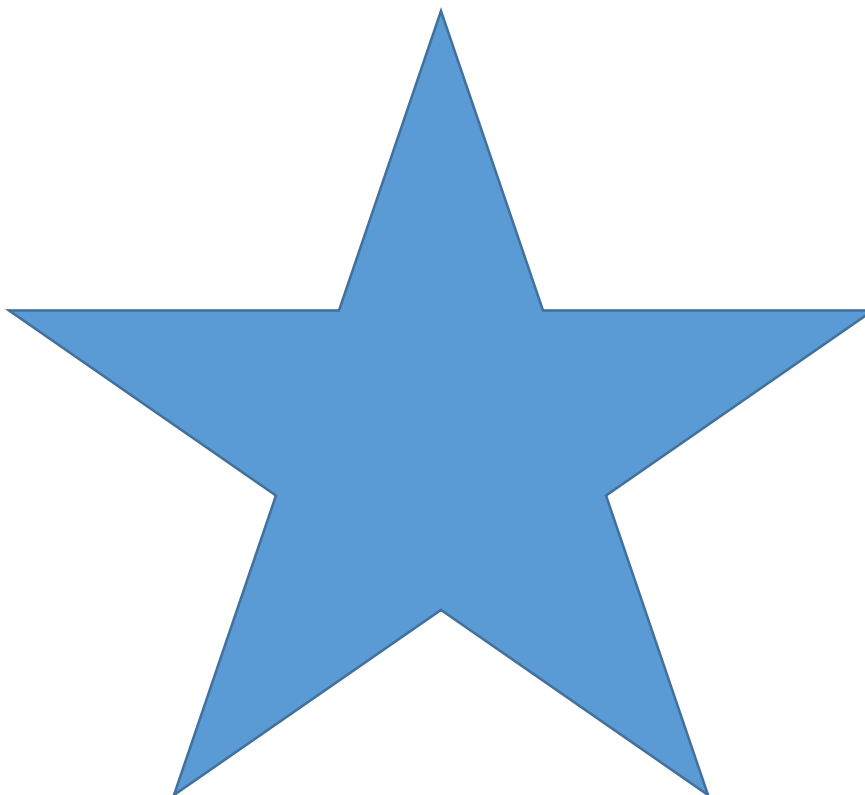
**Synthesis**

**Combustion**

**Decomposition**

**Double Replacement**

**Single Replacement**



# **5 Main Categories**

**Helps us predict things about the reactions**

**Know the reactants?**

**You can predict the products**

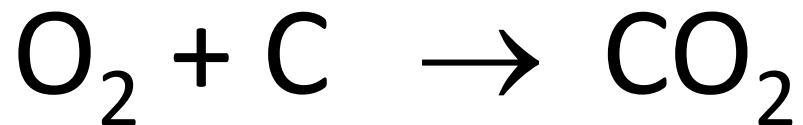
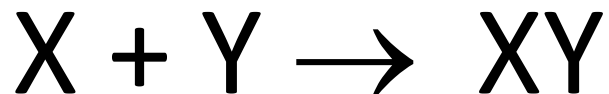
**Know the products?**

**You can predict the reactants**

# Synthesis

*Two things combining into one*

**Example:**



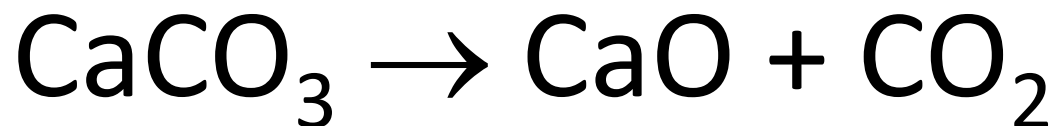
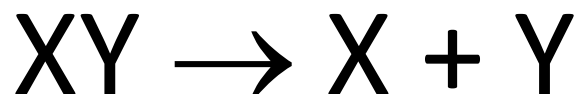
**What to look for:**

More reactants than products

# Decomposition

*One thing falling apart into two*

**Example:**



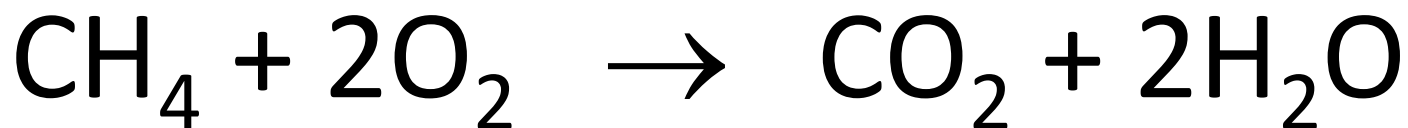
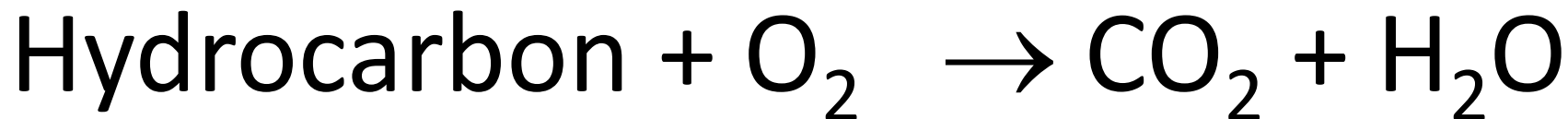
**What to look for:**

More products than reactants

# Combustion

*Burning*

**Example:**



**What to look for:**

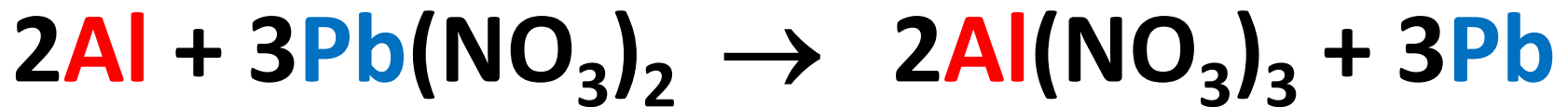
Reactants = Hydrocarbon and  $\text{O}_2$

Products =  $\text{CO}_2$  and  $\text{H}_2\text{O}$

# Single Replacement

*Swapping one element*

Example:



What to look for:

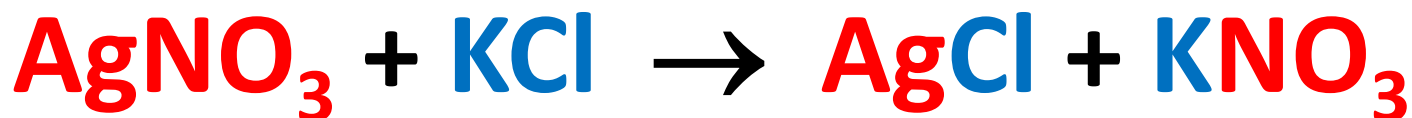
Reactants = 1 element and 1 compound

Products = 1 element and 1 compound,  
but different ones

# Double Replacement

*Swapping two elements*

Example:



What to look for:

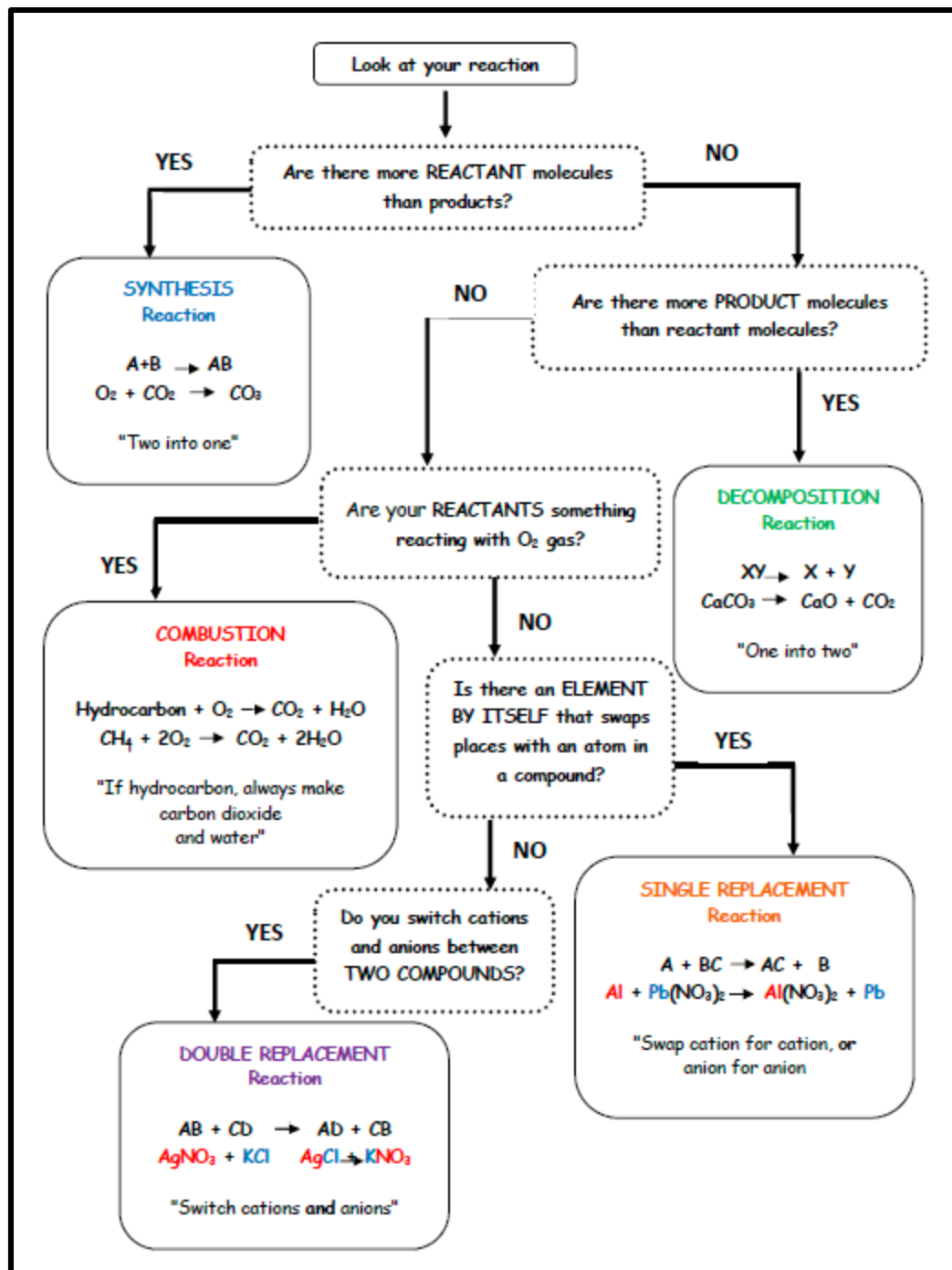
Reactants = 2 Compounds

Products = 2 Compounds but different ones

# For Replacement Rxns

- *If element is a cation, replace it with the other cation. If it is an anion, replace it with the other anion*
- *All neutral compounds need to have a cation and anion when finished (IN THAT ORDER)*
- *You need NEW subscripts – cross over FROM SCRATCH*
- *Careful about diatomic elements in single replacements – they need to be diatomic!*





**YouTube Link to Presentation:**

**<https://youtu.be/RgKtc2vJ5cM>**