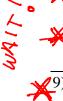
Chalk							
Write your balanced equation:	1	📄 🗋 Prelab Ca	nlculations: <i>show</i>	all work!!!!			
KNaz CO3 + Callz	->/Cal	103 + 2N	acl salt				
You will do the lab starting with 0.018 mol of CaCl <sub>2</sub> < A							
1) How many grams of CaCl <sub>2</sub>	will you start	with?					
Pathway:	1	g A					
molA -> 9 A	molA			g cally			
			2				
molar mess A		(mo) A					
2) If you start with 0.018 mol of CaCl <sub>2</sub> , how many moles of Na <sub>2</sub> CO <sub>3</sub> do you need?							
Pathway:	$\mathfrak{S}$	B					
MoiA > mol B 0.0181	nola	mol B	-	MOI			
mole ratio		mol A	/	Naz CO3			
3) If you start with 0.018 mol of CaCl <sub>2</sub> , how many grams of Na <sub>2</sub> CO <sub>3</sub> do you need? $\sqrt{1}$							
Pathway:	Æ,		₿ <u> </u>				
mai A Thei B > gB 0.01	8 mol A	molB	98 4	9 1			
mr m.mB	T	molALI	MOLB	Na2CO3			
4) If you start with 0.018 mole	es of CaCl <sub>2</sub> ho	w many moles of Ca	aCO3 should you make	?			
Pathway:		•	•				

- 5) If you start with 0.018 mol of CaCl<sub>2</sub>, how many grams of CaCO<sub>3</sub> should you make? *Pathway:*
- 6) If you start with 0.018 moles of CaCl<sub>2</sub> how many moles of NaCl should you make? *Pathway:*
- 7) If you start with 0.018 moles of CaCl<sub>2</sub> how many grams of NaCl should you make? *Pathway:*
- 8) Summarize your calculations in the table below. Use the solubility table in the classroom to fill in the last two rows of the chart.

_		CaCl2	Na2CO3	CaCO <sub>3</sub>	NaCl
L, Y,	# moles	0.018 mol			
	# grams				
<b>*</b>	Solubility				
	Phase				



00 THX PA

Rewrite your balanced equation including the phases.