Calculate how many moles are in the following masses:

- 1) 25 g of NaCl
- 2) 125 g of H₂SO₄
- 3) 100 g of KMnO₄
- 4) 74 g of KCl
- 5) 35 g of CuSO₄(H₂O)₅

Calculate the mass (in g) of the following #of moles:

- 6) 2.5 mol of NaCl
- 7) 0.5 mole of H₂SO₄
- 8) 1.7 mol of KMnO₄
- 9) 0.25 mol of KCl
- 10) 3.2 mol of CuSO₄(H₂O)₅

Calculate how many atoms are in the following number of moles. Put your answer in scientific notation:

- 11) 2 moles
- 12) 1.5 moles
- 13) 0.75 moles
- 14) 15 moles
- 15) 0.35 moles

How many moles are in the following # of atoms.

- 16) 6.02 x 10²³
- 17) 1.204 x 10²⁴
- 18) 1.5 x 10²⁰
- 19) 3.4×10^{26}
- 20) 7.5 x 10¹⁹

Calculate how many moles are in the following masses:

- 1) 25 g of NaCl
- 2) 125 g of H_2SO_4
- 3) 100 g of KMnO₄
- 4) 74 g of KCl
- 5) 35 g of CuSO₄(H₂O)₅

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- 14) 15 moles
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How many moles are in the following # of atoms.

- 16) 6.02×10^{23}
- 17) 1.204 x 10²⁴
- 18) 1.5 x 10²⁰
- 19) 3.4 x 10²⁶ 20) 7.5 x 10¹⁹

Calculate how many moles are in the following masses:

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- 2) 125 g of H₂SO₄3) 100 g of KMnO₄
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- 7) 0.5 mole of H₂SO₄ 8) 1.7 mol of KMnO₄
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10) 3.2 mol of CuSO₄(H₂O)₅ Calculate how many atoms are in the following number of moles. Put your answer in

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