

$$\frac{X \text{ atoms}}{1 \text{ molecule}}$$

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$$\frac{X \text{ grams}}{1 \text{ mole}}$$

Molar Mass

$$\frac{X \text{ grams}}{1 \text{ mole}}$$

Molar Mass

$$\frac{6.02 \times 10^{23} \text{ particles}}{1 \text{ mole}}$$

Avogadro's #

$$\frac{X \text{ grams}}{1 \text{ L}}$$

Density

$$\frac{X \text{ grams}}{1 \text{ L}}$$

Density

A = what you have/know
B = what you want

$$\frac{6.02 \times 10^{23} \text{ particles}}{1 \text{ mole}}$$

Avogadro's #

$$\frac{\text{Coefficient of B}}{\text{Coefficient of A}}$$

Mole Ratio

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