|  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- | --- |
|

|  |  |
| --- | --- |
| Dihydrogen monoxide | *Formula* |
| Molar Mass |  |
| # of Grams | 54g |
| # of Moles |  |

 | Balance the following Reaction\_\_\_ N2 + \_\_\_H2 🡪 \_\_\_NH3What type of reaction is this? |

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
|  | **2Al** | **+** | **3Cu(SO4)** | **🡪** | **Al2(SO4)3** | **+** | **3Cu** |
| **Molar Mass** |  |  |  |  |  |  |  |
| **# of Moles** |  |  |  |  |  |  |  |
| **# of grams** |  |  |  |  | 100g |  |  |

 |
| Predict Products and then Balance the following Reaction\_\_\_ NaCl + \_\_\_F2 🡪 What type of reaction is this? |

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
|  | **2C2H6** | **+** | **7O2** | **🡪** | **4CO2** | **+** | **6H2O** |
| **Molar Mass** |  |  |  |  |  |  |  |
| **# of Moles** |  |  |  |  |  |  |  |
| **# of grams** | 12g |  |  |  |  |  |  |

 |

|  |  |
| --- | --- |
| Oxygengas | *Formula* |
| Molar Mass |  |
| # of Grams | 96g |
| # of Moles |  |

 |
|

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
|  | **2Mg** | **+** | **O2** | **🡪** | **2MgO** |
| **Molar Mass** |  |  |  |  |  |
| **# of Moles** |  |  |  |  |  |
| **# of grams** |  |  | 20g |  |  |

 |

|  |  |
| --- | --- |
| Carbon dioxide | *Formula* |
| Molar Mass |  |
| # of Grams |  |
| # of Moles | 0.5 mol |

 | Write the formulas for the skeleton equation and then balance. *Potassium Chlorate decomposes into  Potassium Chloride and Oxygen Gas* |
|

|  |  |
| --- | --- |
| Cobalt (III)Cyanide | *Formula* |
| Molar Mass |  |
| # of Grams | 50g |
| # of Moles |  |

 | Balance the following Reaction\_\_\_ O2 + \_\_\_H2 🡪 \_\_\_H2OWhat type of reaction is this? |

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
|  | **2Fe** | **+** | **SnCl4** | **🡪** | **2FeCl2** | **+** | **Sn** |
| **Molar Mass** |  |  |  |  |  |  |  |
| **# of Moles** |  |  |  |  |  |  |  |
| **# of grams** |  |  |  |  | 10g |  |  |

 |
| Balance the following Reaction\_\_\_ H2CO3 🡪 \_\_\_H2O + \_\_\_CO2What type of reaction is this? |

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
|  | **2C6H14** | **+** | **19O2** | **🡪** | **12CO2** | **+** | **14H2O** |
| **Molar Mass** |  |  |  |  |  |  |  |
| **# of Moles** |  |  |  |  |  |  |  |
| **# of grams** | 8g |  |  |  |  |  |  |

 |

|  |  |
| --- | --- |
| Sulfur trioxide | *Formula* |
| Molar Mass |  |
| # of Grams | 30g |
| # of Moles |  |

 |
|

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
|  | **P4** | **+** | **5O2** | **🡪** | **2P4O10** |
| **Molar Mass** |  |  |  |  |  |
| **# of Moles** |  |  | 0.5mol |  |  |
| **# of grams** |  |  |  |  |  |

 |

|  |  |
| --- | --- |
| Aluminum chloride | *Formula* |
| Molar Mass |  |
| # of Grams |  |
| # of Moles | 0.5 mol |

 | Write the formulas for the skeleton equation and then balance. *Tricarbon octahydride combusts* |
|

|  |  |
| --- | --- |
| Titanium (IV)oxide | *Formula* |
| Molar Mass |  |
| # of Grams | 17g |
| # of Moles |  |

 | Balance the following Reaction\_\_\_ O2 + \_\_\_S8 🡪 \_\_\_SO3What type of reaction is this? |

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
|  | **2Cl2** | **+** | **CH4** | **🡪** | **CCl4** | **+** | **2H2** |
| **Molar Mass** |  |  |  |  |  |  |  |
| **# of Moles** |  |  |  |  |  |  |  |
| **# of grams** |  |  |  |  | 10g |  |  |

 |
| Predict Products and then Balance the following Reaction\_\_\_ AlBr3 + \_\_\_K2SO4 🡪 What type of reaction is this? |

|  |  |  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- | --- | --- |
|  | **\_\_C4H10** | **+** | **\_\_O2** | **🡪** | **\_\_CO2** | **+** | **\_\_H2O** |
| **Molar Mass** |  |  |  |  |  |  |  |
| **# of Moles** |  |  |  |  |  |  |  |
| **# of grams** | 12g |  |  |  |  |  |  |

 |

|  |  |
| --- | --- |
| Bariumhydroxide | *Formula* |
| Molar Mass |  |
| # of Grams |  |
| # of Moles | 0.75 mol |

 |
|

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
|  | **\_\_\_P4** | **+** | **\_\_\_Cl2** | **🡪** | **\_\_\_PCl5** |
| **Molar Mass** |  |  |  |  |  |
| **# of Moles** |  |  | 0.5mol |  |  |
| **# of grams** |  |  |  |  |  |

 |

|  |  |
| --- | --- |
| Oxygen difluoride | *Formula* |
| Grams |  |
| Molar Mass |  |
| # Moles | 2.5 mol |

 | Write the formulas for the skeleton equation and then balance. *Sodium Bicarbonate decomposes into Sodium Carbonate, Water, and Carbon Dioxide* |