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| How much energy does it take to increase the temperature of ice at 0°C to water at 75°C? |  | If ice at -1°C is heated to 101°C how much energy is released/absorbed?  |
| If 11g of ice is heated to 25°C how much energy is absorbed? |  | When steam at 110°C cools to ice at -10°C how much energy is released/absorbed? |
| How much energy is released/absorbed when water at 25°C cools to ice at -5°C? When might this happen in real life? |  | How much energy is released when 20g of water increases from 10°C to 94°C? |
| When 25g of steam at 103°C cools to water at 25°C how much energy is released?  |  | 3 g H2O105°C-5°C |
| When 35g of water at 100°C is boiled away how much energy is absorbed by the water? |  | 7 g H2O-8°C |

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| 105°C10 g H2O0°C |  | 21 g H2O-5°C50°C |
| (g)(l)61 g H2O100°C |  | Csolid =3.5 J/g°CLfus = 185 J/gCliq = 0.95 J/g°CLvap = 1950 J/g85°C26°C20 g substance A4°C |
| 24g H2O0°C100°C |  | Which will heat/ cool faster? Substance X or Substance Y? |
| 100°C5 g H2O0°C |  | Which phase change is shown here?-3°C |
| 15 g H2O101°C-1°C |  | Which phase change isshown here?100°C0°C |