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| How much energy does it  take to increase the temperature of ice at 0°C to water at 75°C? |  | If ice at -1°C is heated to  101°C how much energy is released/absorbed? |
| If 11g of ice is heated to  25°C how much energy is absorbed? |  | When steam at 110°C  cools to ice at -10°C how  much energy is  released/absorbed? |
| How much energy is released/absorbed when  water at 25°C cools to ice  at -5°C? When might this  happen in real life? |  | How much energy is released when 20g of water increases from 10°C to 94°C? |
| When 25g of steam at  103°C cools to water at  25°C how much energy is released? |  | 3 g H2O  105°C  -5°C |
| When 35g of water at 100°C  is boiled away how much  energy is absorbed by  the water? |  | 7 g H2O  -8°C |

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| 105°C  10 g H2O  0°C |  | 21 g H2O  -5°C  50°C |
| (g)  (l)  61 g H2O  100°C |  | Csolid =3.5 J/g°C  Lfus = 185 J/g  Cliq = 0.95 J/g°C  Lvap = 1950 J/g  85°C  26°C  20 g substance A  4°C |
| 24g H2O  0°C  100°C |  | Which will heat/ cool faster? Substance X or Substance Y? |
| 100°C  5 g H2O  0°C |  | Which phase change is  shown here? -3°C |
| 15 g H2O  101°C  -1°C |  | Which phase change is shown here? 100°C  0°C |