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2. Reaction rate refers to how quickly or slowly the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_disappear

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1. What is meant by the term Reaction Rate?
2. What is the collision theory?
3. What is the activation energy?
4. What is a catalyst and why is it different from a reactant in an equation?
5. What are the FOUR major factors that affect reaction rate?
6. How does each of the factors above affect the rate of reaction?
7. Draw an exothermic reaction graph shown with and without a catalyst?
8. Why would a mixture of gases react faster when the volume they occupy is decreased?
9. Why would iron filings rust faster than an iron nail?
10. How would the increasing pressure of reactive components of a gaseous mixture affect the rate at which the components react with one another?
11. How would you change temperature of a reaction if you wanted to increase the rate of reaction? Explain how this effects the reaction using the collision theory.
12. Why would the rate of reaction decrease as the reaction produces more products?
13. If you put 100g of NaOH in cube form and 200g of NaOH in powdered form which will react with HCl at a faster rate and list the reasons why?
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