

Task E: - Review the data from an experiment below and answer the questions:

Let's see how reaction rates are affected by changes in concentration, temperature, and activation energy. The results of an experiment are listed below.

Reaction Rates

Trial	Reaction Type	[A] mol/dm ³	Temperature K	Activation Energy kJ/mol	Reaction Rate mol/dm ³ ·s
1	Rate = k[A][A]	0.10	298	65	+ 0.0004
2	Rate = k[A][A]	0.25	298	65	+ 0.0025
3	Rate = k[A][A]	0.50	298	65	+ 0.0099
4	Rate = k[A][A]	0.50	250	65	+ 0.0001
5	Rate = k[A][A]	0.50	373	65	+ 0.3990
6	Rate = k[A][A]	0.50	373	50	+ 0.4990
7	Rate = k[A][A]	0.50	373	80	+ 0.0152