* + - Strong Acid Weaker Conjugate Base   
       *(not much effect on pH)*
    - Weak Acid Stronger Conjugate Base  
       *(potential effect on pH)*
    - Strong Base Weaker Conjugate Acid  
       *(not much effect on pH)*
    - Weak Base Stronger Conjugate Acid

*(potential effect on pH)*

* + - Ion from a Strong Acid Neutral   
      (*is a weak conj. base*)
    - Ion from a Weak Acid Basic   
      (*is a strong conj. base*)
    - Ion from a Strong Base Neutral   
      (*is a weak conj. acid*)
    - Ion from a Weak Base Acidic   
      (*is a strong conj. acid*)
    - Cation is a charged metal ion, and anion is from a strong acid Acidic metal hydrate + Neutral anion - salt is acidic

* + - Neutral + Acidic = Acidic
    - Neutral + Basic = Basic
    - Neutral + Neutral = Neutral
    - Acidic + Basic = ?   
      *Use Ka and Kb to determine* Ka > Kb 🡪 Acidic

Ka < Kb 🡪 Basic  
Ka = Kb 🡪 Neutral

* + - Kw = Ka x Kb Kw = 1.0 x 10-14 (*if at 25 °C, may be different if not at 25°C*)

If you are looking for the Ka of an acidic conjugate ion, use Kw and the Kb of the base it came from

If you are looking for the Kb of a basic conjugate ion, use Kw and the Ka of the acid it came from

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