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| **Galvanic/Voltaic** | **Electrolytic** |
| 1. Chemical → Electric energy
 | 1. Electric → Chemical energy
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| 1. Spontaneous redox reaction - produces electric current
 | 1. Not spontaneous - requires electric supply
 |
| 1. Two half cells are separated- connected through salt bridge
 | 1. Both electrodes can be put in same container
 |
| 1. Anode (-) Cathode (+)Oxidation = AnodeReduction = Cathode“An Ox, Red Cat”
 | 1. Anode (+) Cathode (-)Oxidation = AnodeReduction = CathodeStill “An Ox, Red Cat” you just change the  algebraic sign
 |
| 1. e- supplied by species being oxidized- move from anode to cathode
 | 1. Power source (ex. Battery) supplies the e- - go in through the cathode and come out  through the anode
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| 1. ∆G < 0, K > 1
 | 1. ∆G > 0, K < 1
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**R-46**