**Dougherty Valley HS AP Chemistry**

**S-81**

**Acid Base Reactions**

**Quick Check #6**

**Name: Date: Period: Seat #:**

🞎 **pH of a weak acid solution**

Calculate the pH of a 0.200 M solution of nitrous acid, HNO2. Ka of HNO2 = 4.0 x 104.

🞎 **Salt solutions**

A solution of NaNO2 will be \_\_\_\_\_\_\_\_\_\_\_ (acidic, basic, neutral).

Write the ***net*** equation for the equilibrium involved when NaNO2 dissolves in water.

Write the equilibrium expression for the above equation. Should this be labeled Kc, Ka, Kb, Keq?

Calculate the pH of a 0.100 M solution of NaNO2.

🞎 **Acid-Base Neutralization**

Write the balanced net equation for:

A solution of sulfurous acid is added to a suspension of magnesium hydroxide