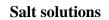
Dougherty Valley HS AP Chemistry Acid Base Reactions Quick Check #6		S-81		
Name:	Date:	Period:	Seat #:	

pH of a weak acid solution

Calculate the pH of a 0.200 M solution of nitrous acid, HNO₂. K_a of HNO₂ = 4.0 x 10⁻⁴.



A solution of NaNO₂ will be _____ (acidic, basic, neutral).

Write the *net* equation for the equilibrium involved when NaNO₂ dissolves in water.

Write the equilibrium expression for the above equation. Should this be labeled K_c, K_a, K_b, K_{eq}?

Calculate the pH of a 0.100 M solution of NaNO₂.

Acid-Base Neutralization

Write the balanced net equation for: A solution of sulfurous acid is added to a suspension of magnesium hydroxide