Dougherty V Acid Base Ed Quick Check	7alley HS AP Chemi quilibrium x #1	istry				S-76
Name:			Date:	Per	iod: Seat	# :
Acid - Consider the Identify the	Base Equilibria ne following equil e two conjugate a	ibrium: HCN -	+ $H_2O(l) \leftrightarrows$	$H_3O^+ + CN^-$	K _a for HC	$N = 4.0 \times 10^{-10}$
Draw a box	around the more	appropriate repre	esentation of th	is equilibrium:		
HC	$CN + H_2C$	$\mathbf{D}(l)$ \Le H_3O^+	+ CN ⁻	$HCN + H_2O(l)$	\Rightarrow H ₃ O ⁺	$+ CN^{-}$
Which acid	l is weaker?					
🗌 Conju	igate Acids and I	Bases				
Wh	at is the conjugate <i>bas</i>	$e of H_2O?$				
	conjugate aci	d of NH ₃ ?				
	conjugate bas	e of OH ⁻ ?				
	conjugate <i>aci</i>	<i>d</i> of HCO ₃ ⁻ ?				
рнс а	alculations					
Fill	in the chart below $[\mathbf{H}_{3}\mathbf{O}^{+}]$	<u>v:</u>	nH	nOH	Acidic or Resid	•
	2.0×10^{-3}		PII	pon	ACIUIC OF DASI	
			6.25			
		5.6 x 10 ⁻²				