**Dougherty Valley HS AP Chemistry**

**S-88**

**Electrochemistry**

**Quick Check #1**

**Name: Date: Period: Seat #:**

🞎 **Electrolysis Reactions**

Use your Reduction Potential Chart to determine the reaction at the anode, reaction at the cathode, and the overall reaction during the electrolysis of a solution of copper(II) chloride.

🞎 **Electrolysis Stoichiometry**

A current of 2.50 amps is passed through a solution of Ni(NO3)2 for 2.00 hours.

What mass of Ni metal is deposited?

🞎 **Electrochemical Cell**

Sketch the cell made from Au in a 1.0 M solution of gold(III) nitrate and Zn in a 1.0 M solution

of Zn(II) sulfate.

Au3+ + 3e- → Au(s) +1.50

Zn2+(aq) + 2e- → Zn(s) -0.763

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| --- | --- | --- |
| Anode Reaction | Sketch the Cell | Cathode Reaction |
|  |  |  |
|  | Overall Reaction & E° (volts) |  |
|  | | |