Name: Date: Period: Seat #:
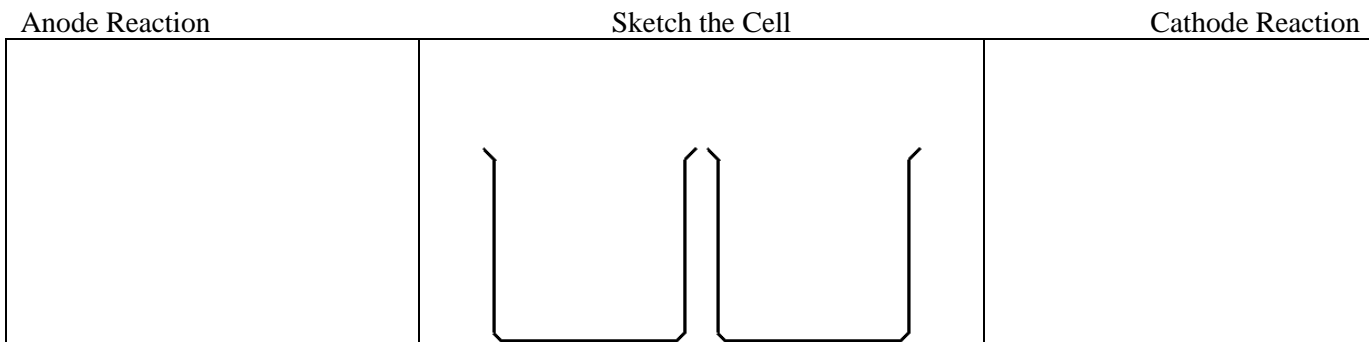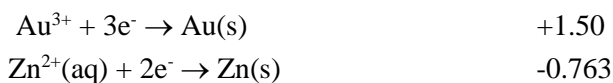
☐ **Electrolysis Reactions**

Use your Reduction Potential Chart to determine the reaction at the anode, reaction at the cathode, and the overall reaction during the electrolysis of a solution of copper(II) chloride.

☐ **Electrolysis Stoichiometry**

A current of 2.50 amps is passed through a solution of $Ni(NO_3)_2$ for 2.00 hours.
What mass of Ni metal is deposited?

☐ **Electrochemical Cell**

Sketch the cell made from Au in a 1.0 $\underline{M}$ solution of gold(III) nitrate and Zn in a 1.0 $\underline{M}$ solution of Zn(II) sulfate.

$$Au^{3+} + 3e^- \rightarrow Au(s) \qquad +1.50$$
$$Zn^{2+}(aq) + 2e^- \rightarrow Zn(s) \qquad -0.763$$

Anode Reaction   Sketch the Cell   Cathode Reaction

Overall Reaction & E° (volts)