**No Calc Solution Stoich Fun**

1. **NaOH + HCl → H2O + NaCl(aq)**

**40.0mL of 0.125M NaOH is used to titrate 5.0mL of HCl. What is the concentration of the acid?**

1. **30.0mL of 0.10M NaCl is mixed with 20.0mL 0.20M AlCl3. Find [Cl-].**
2. **20.0g NaOH is dissolved in 250.0mL solution. It is then added to enough water to make 1.5L of NaOH. Find [NaOH] of the resulting solution.**
3. **108.0g of C5H12 (molar mass 72.15g mol-1) is dissolved in enough water to make 750.0mL of solution. What is the concentration?**
4. **Acetic acid, (HC2H3O2 ) is mixed with sodium hydroxide to neutralize. Write the NIE.**
5. **Is the following reaction redox? Justify. C3H6 + O2 → CO2 + H2O**
6. **Draw a particle model showing the products. Pb(NO3)2 + CaCl2 → PbCl2(s) + Ca(NO3)2.**

**7b. Describe how the particle model differs (if at all) if there is excess CaCl2.**