**Ksp Mental Math**

1. The ionic salt, XY, is added to water and the molar solubility is 1.2 x 10-3. Find Ksp.
2. The ionic salt XY2 has a Ksp = 3.2 x 10-5. Find the molar solubility.
3. Ksp = 2.1 x 10-6 for the ionic salt XY. XY is added to a solution that already has 0.3 M NaY added. Find the molar solubility of XY.
4. Determine if the ionic salt, XY2, will precipitate out from mixing 50mL of 0.3M X2+(aq) and 25mL of 0.3M Y-(aq) together. Ksp = 3 x 10-3.

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