1. What is the pH of a 1.0 x 10-2 molar solution of HCN? (Ka = 4.0 x 10-10)?
2. 10
3. Between 7 and 10
4. 7
5. Between 4 and 7
6. 4

2. What is the H+(aq) concentration in 0.05 M HCN(aq)? (Ka for HCN is 5.0 x 10-10)?

3. The pH of a 0.01 M HNO2 (aq) solution is in which of the following ranges? (Ka = 4.0 x 10-4)?

1. Between 1 and 2
2. Between 2 and 3
3. Between 4 and 5
4. Between 6 and 7
5. What is the pH of a 1.0 x 10-2 molar solution of HCN? (Ka = 4.0 x 10-10)?
6. 10
7. Between 7 and 10
8. 7
9. Between 4 and 7
10. 4

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