**No Calc Equilibrium**

| 3A(aq) + 2B(s) ↔ 2 C(aq) + D(g) | 1. Write the expression for the equilibrium constant.
2. If [A] = 0.20M, [C] = 0.40M and [D] = 0.10M, evaluate the value for K.
3. Evaluate Q if each concentration is 0.20M
 |
| --- | --- |
| 2A(g) ↔ 2 B(g) + C(g) | 1. Write the expression for the equilibrium constant.
2. Evaluate Q if the pressure of A = 1.0atm, B = 0.4atm and C = 0.3atm.
 |
| A(aq) + ½ B(aq) ↔ 4 C(aq) + D(s) | 1. Write the expression for the equilibrium constant.
2. If [A] = 0.01M, [B] = 0.16M and [C] = 0.20M, evaluate the value for K.
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